

CH #9
SECTION #1

Assessment

Stoichiometry

Quiz CH #9 SECTION #1

Section Quiz: Introduction to Stoichiometry

In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question.

C

1. Stoichiometry is the branch of chemistry that deals with elements in compounds and with reactants and products in chemical reactions, focusing on

- a. bonding.
- b. energy transfers.
- c. mass relationships.
- d. physical characteristics.

A

2. Reaction stoichiometry is based on chemical equations and

- a. the law of conservation of mass.
- b. the ideal gas law.
- c. the law of conservation of momentum.
- d. Avogadro's number.

C

3. The number of significant figures in an answer to a stoichiometry problem is determined only by

- a. the number of significant figures in the mole ratio used to solve the problem.
- b. the number of decimal places in the molar masses of substances in the chemical equation.
- c. the number of significant figures of any measured quantities in the problem.
- d. the number of significant figures in the molar masses of substances in the chemical equation.

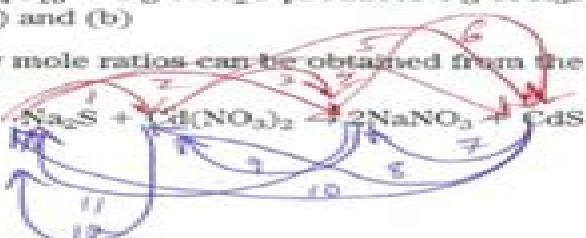
D

4. The chemical equation $P_4O_{10} + 6H_2O \rightarrow 4H_3PO_4$ can be interpreted correctly as

- a. 1 molecule of P_4O_{10} + 6 molecules of H_2O form 4 molecules of H_3PO_4 .
- b. 1 mol of P_4O_{10} + 6 mol of H_2O reacts to form 4 mol of H_3PO_4 .
- c. 1 g of P_4O_{10} + 6 g of H_2O produces 4 g of H_3PO_4 .
- d. Both (a) and (b)

C

5. How many mole ratios can be obtained from the following chemical equation?



- a. 1
- b. 4
- c. 12
- d. 16

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