

**Question 1**

(a) Using only the periodic table, arrange each set of atoms in order of increasing radius:

(1) Ba, Ca, Na

Na < Ca < Ba

(2) In, Sn, As

As < Sn < In

(3) Cs, K, Al

Al < K < Cs

(b) Which neutral atom is isoelectronic with each of the following ions?

(1) H<sup>+</sup>

None

(The H<sup>+</sup> ion has 0 electrons. It is impossible for a neutral atom to have 0 electrons.)

(2) Zr<sup>4+</sup>

Kr

(3) Mn<sup>7-</sup>

Ar

(4) I<sup>-</sup>

Xe

(c) Explain the following variations in atomic or ionic radii:

(1) I<sup>-</sup> > I > I<sup>+</sup>

As the atomic number stays constant and the number of electrons increases, the electron-electron repulsions increase. This leads to the electrons spreading out, and the anion becomes larger. The reverse is also true, so the cation is smaller.

(2) Ca<sup>2+</sup> > Mg<sup>2+</sup> > Be<sup>2+</sup>

For cations with the same charge, ionic radii increase going down a column because of the increase in principal quantum number, which means the outer electrons are farther away from the nucleus on average. This leads to larger ion size.

# Ionic Radius Answer Key

**Joseph Brennan**

## Ionic Radius Answer Key:

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