

Reminders:

- In a neutral atom the number of protons equals the number of electrons.
- An atom can NEVER gain or lose protons.
- The number of protons equals the atomic number.

Important! NAME: Answer Key

Ion Practice Set

- What is an ion? **a charged ion - it can be positively charged (cation) or negatively charged (anion)**
- What does the number next to the ions signify? **the charge of the ion**

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED
ex	Fluorine	F ⁻	9	10	gained one
1	Iodine	I ⁻	53	54	gained one
2	Sulfur	S ²⁻	16	16+2 = <u>18</u>	gained two
3	Potassium	K ⁺	19	19-1 = <u>18</u>	lost one
4	Calcium	Ca ²⁺	20	20-2 = <u>18</u>	lost two
5	Bromine	Br ⁻	35	36	gained one
6	Strontium	Sr ²⁺	38	38-2 = <u>36</u>	lost two
7	Hydrogen	H ⁺	1	1-1 = <u>0</u>	lost one
8	Oxygen	O ²⁻	8	8+2 = <u>10</u>	gained two
9	Magnesium	Mg ²⁺	12	12-2 = <u>10</u>	lost two
10	aluminum	Al ³⁺	13	10	lost three
11	Selenium	Se ²⁻	34	36	gained two
12	Hydrogen	H ⁺	1	1+1 = <u>2</u>	gained one
13	Lithium	Li ⁺	3	3-1 = <u>2</u>	lost one
14	Rubidium	Rb ⁺	37	37-1 = <u>36</u>	lost one
15	Chlorine	Cl ⁻	17	18	gained one

this means it has an extra electron

this means it has 2 more protons than electrons so 2 electrons were lost!

[**Ion And Isotopes Practice Set Answers**](#)

Thomas Griffiths

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