

KEY

## Ionic Formula Practice

Name	Cation	Anion	Formula
1. calcium bromide	$\text{Ca}^{+2}$	$\text{Br}^{-1}$	$\text{Ca Br}_2$
2. magnesium phosphide	$\text{Mg}^{+2}$	$\text{P}^{-3}$	$\text{Mg}_3 \text{P}_2$
3. aluminum oxide	$\text{Al}^{+3}$	$\text{O}^{-2}$	$\text{Al}_2 \text{O}_3$
4. barium sulfide	$\text{Ba}^{+2}$	$\text{S}^{-2}$	$\text{Ba S}$
5. sodium phosphide	$\text{Na}^{+1}$	$\text{P}^{-3}$	$\text{Na}_3 \text{P}$
6. strontium chloride	$\text{Sr}^{+2}$	$\text{Cl}^{-1}$	$\text{Sr Cl}_2$
7. calcium nitride	$\text{Ca}^{+2}$	$\text{N}^{-3}$	$\text{Ca}_3 \text{N}_2$
8. ammonium bromide	$\text{NH}_4^{+1}$	$\text{Br}^{-1}$	$\text{NH}_4 \text{Br}$
9. calcium sulfate	$\text{Ca}^{+2}$	$\text{SO}_4^{-2}$	$\text{Ca SO}_4$
10. magnesium nitrate	$\text{Mg}^{+2}$	$\text{NO}_3^{-1}$	$\text{Mg}(\text{NO}_3)_2$
11. sodium carbonate	$\text{Na}^{+1}$	$\text{CO}_3^{-2}$	$\text{Na}_2 \text{CO}_3$
12. zinc hydroxide	$\text{Zn}^{+2}$	$\text{OH}^{-1}$	$\text{Zn}(\text{OH})_2$
13. barium nitrite	$\text{Ba}^{+2}$	$\text{NO}_2^{-1}$	$\text{Ba}(\text{NO}_2)_2$
14. aluminum sulfite	$\text{Al}^{+3}$	$\text{SO}_3^{-2}$	$\text{Al}_2(\text{SO}_3)_3$
15. lead II nitrate	$\text{Pb}^{+2}$	$\text{NO}_3^{-1}$	$\text{Pb}(\text{NO}_3)_2$
16. copper I phosphate	$\text{Cu}^{+1}$	$\text{PO}_4^{-3}$	$\text{Cu}_3 \text{PO}_4$
17. ammonium carbonate	$\text{NH}_4^{+1}$	$\text{CO}_3^{-2}$	$(\text{NH}_4)_2 \text{CO}_3$
18. tin IV nitrate	$\text{Sn}^{+4}$	$\text{NO}_3^{-1}$	$\text{Sn}(\text{NO}_3)_4$
19. nickel II acetate	$\text{Ni}^{+2}$	$\text{C}_2\text{H}_3\text{O}_2^{-1}$	$\text{Ni}(\text{C}_2\text{H}_3\text{O}_2)_2$
20. chromium III sulfate	$\text{Cr}^{+3}$	$\text{SO}_4^{-2}$	$\text{Cr}_2(\text{SO}_4)_3$
21. iron III nitrite	$\text{Fe}^{+3}$	$\text{NO}_2^{-1}$	$\text{Fe}(\text{NO}_2)_3$
22. cobalt II hydroxide	$\text{Co}^{+2}$	$\text{OH}^{-1}$	$\text{Co}(\text{OH})_2$

# Ionic Formula Writing Kit Lab Answer Key

**Tracy Lynn Haroff**



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