

20. Which of the following correctly represents the ionization of an atom?

- a)  $\text{Cl(g)} + e^- \rightarrow \text{Cl}^-(\text{g})$   
**b)  $\text{Na(g)} \rightarrow \text{Na}^+(\text{g}) + e^-$**   
 c)  $\text{Na(s)} - e^- \rightarrow \text{Na}^+(\text{g})$   
 d)  $\text{Cl}_2(\text{g}) \rightarrow 2 \text{Cl(g)}$

21. Which of the following is likely to have the largest atomic radius?

- ~~a) H~~ ~~b) Mn~~ ~~c) Cl~~ **d) Rb** ~~e) Ag~~  
 MOST LOW  $\hat{=}$  LEFT

22. Which one of the following isoelectronic species has the smallest radius? *most protons*

- a)  $\text{Mg}^{2+}$  12  $p^+$**  d)  $\text{F}^-$  9  $p^+$   
 b)  $\text{Na}^+$  11  $p^+$  e)  $\text{O}^{2-}$  8  $p^+$   
 c) Ne 10  $p^+$

NOTE: explain your reasoning on the last page.

23. Which of the following has the greatest ionization energy? *smallest*

- a) K b) Ca c) Fe d) Ga **e) Br**  
 all period 4

24. Which of the following has the lowest ionization energy? *LARGEST*

- a) Li b) Na c) K d) Rb **e) Cs**  
 same family

25. The successive ionization energies for one of the period three elements is listed below. Which element is referred to?

$E_1$	577.4 kJ/mol
$E_2$	1,816 kJ/mol
$E_3$	2,744 kJ/mol
$E_4$	11,580 kJ/mol
$E_5$	15,030 kJ/mol

} valence electrons  
 } ~~core electrons~~

- a) Na ~~b) Mg~~ **c) Al** d) Si e) P

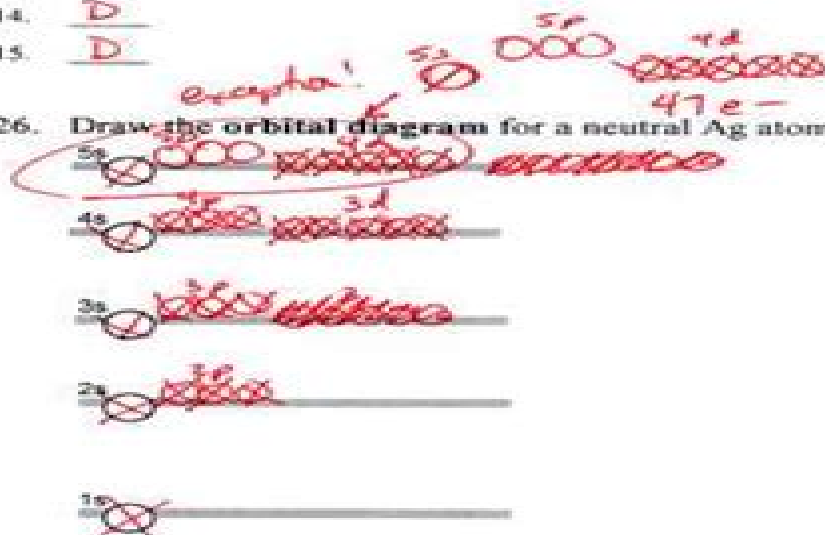
NOTE: explain your reasoning on the last page.

# Multiple Choice Answers:

Use CAPITAL LETTERS for your answers....

- |              |              |
|--------------|--------------|
| 1. <u>B</u>  | 16. <u>E</u> |
| 2. <u>D</u>  | 17. <u>A</u> |
| 3. <u>C</u>  | 18. <u>A</u> |
| 4. <u>D</u>  | 19. <u>A</u> |
| 5. <u>C</u>  | 20. <u>B</u> |
| 6. <u>E</u>  | 21. <u>D</u> |
| 7. <u>A</u>  | 22. <u>A</u> |
| 8. <u>C</u>  | 23. <u>E</u> |
| 9. <u>B</u>  | 24. <u>E</u> |
| 10. <u>B</u> | 25. <u>C</u> |
| 11. <u>D</u> |              |
| 12. <u>B</u> |              |
| 13. <u>C</u> |              |
| 14. <u>D</u> |              |
| 15. <u>D</u> |              |

26. Draw the orbital diagram for a neutral Ag atom: *excepta! 47 e-*



Write the electron configuration for silver:

27. Long form:  
 $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^1$   
 28. Short form:  
 $[\text{Kr}] 4d^{10} 5s^1$

# Ionic And Metallic Bonding Answer Key

**John T. Moore, Richard H. Langley**



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