

# Chemistry: Ionic & Covalent Bonds Unit Study Guide

## Covalent Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Covalent Compounds.

- Bond (Ionic / Covalent)
- Charge (+) / (-)
- Atoms (3 / 2)
- (High / Low) Conductivity
- (Solid / Liquid / Gas) State
- Polarity (Yes / No)
- Nonpolar

2. Define the following numbers

State	Number
Solid	1
Liquid	2
Gaseous	3
Gas	4

3. What are the two types of bonds?

4. Complete the following table.

FORMULA	NAME
HF <sub>2</sub>	Hydrogen Fluoride
Na <sub>2</sub> O	Sodium Oxide
Si <sub>2</sub> S <sub>2</sub>	Silicon Sulfide
Li <sub>2</sub> CO <sub>3</sub>	Lithium Carbonate

5. What does VSEPR stand for?

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Period: \_\_\_\_\_

## Ionic Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Ionic Compounds.

- Bond (between a (Metal / Nonmetal) and (Metal / Nonmetal))
- (Large / Small) Electronegativity Difference
- The (Metal / Nonmetal) form Cations, which (Gains / Loses) Electrons
- The (Metal / Nonmetal) form Anions, which (Gains / Loses) Electrons
- (High / Low) Melting & Boiling Point
- (Conducts / Does not Conduct) Electricity in Solution
- (Solid / Liquid / Gas) at Room Temperature

2. What charges would the following periodic table elements have in a chemical bond?

Period	Charge
Metalloids	
Metals	
Nonmetals	
Halogens	
Neonogens	

3. Write the atomic number that is represented by the following numbers.

Atomic Number	Number
1	Hydrogen
2	Helium
3	Lithium
4	Boron
5	Carbon
6	Nitrogen
7	Phosphorus
8	Sulfur
9	Chlorine
10	Argon

4. Complete the chart below by filling in the missing boxes.

SATION	ANION	FORMULA	NAME
Mg <sup>2+</sup>	Br <sup>-</sup>		
		CuO	
			Iron (III) chloride
		(NH <sub>4</sub> ) <sub>2</sub> PO <sub>4</sub>	

# Ionic And Covalent Bonds Study Guide

**Darrell D. Ebbing**

## Ionic And Covalent Bonds Study Guide:

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