

Electron Configuration Practice

Chemistry

How to write an electron configuration:

- A. Determine the total number of electrons to be represented.
- B. Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- C. The sum of the superscripts should equal the total number of electrons. Example:
 $_{12}Mg \ 1s^2 \ 2s^2 \ 2p^6 \ 3s^2$

I. Configuration Writing Practice

A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the pr per number of electrons for the element) are filled.

1. Na: $1s^2 \ 2s^2 \ 2p^6 \ 3s^1$
2. Pb: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^{14} \ 5d^{10} \ 6p^2$
3. Sr: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2$
4. U: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^{14} \ 5d^{10} \ 6p^6 \ 7s^2 \ 5f^3 \ 6d^1$
5. N: $1s^2 \ 2s^2 \ 2p^3$
6. Ag: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^1 \ 4d^{10}$
7. Ti: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^2$
8. Ce: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^1 \ 5d^1$
9. Cl: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^5$
10. Hg: $1s^2 \ 2s^2 \ 2p^6 \ 3s^2 \ 3p^6 \ 4s^2 \ 3d^{10} \ 4p^6 \ 5s^2 \ 4d^{10} \ 5p^6 \ 6s^2 \ 4f^{14} \ 5d^{10}$

Electron Configuration Packet Answers

Morris Hein, Susan Arena

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Foundations of College Chemistry Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals. For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed. It provides them with clear and logical explanations of chemical concepts and problem solving. They'll learn how to apply concepts with the help of worked out examples. In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis.

Basic Chemistry Steven S. Zumdahl, 2004 Description Not Yet Available

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Modern Physics Robert L. Sproull, 1956

Essentials of Chemistry Ralph A. Burns, John William Hill, 1995 The second edition of this chemistry textbook that uses practice examples and applications relating chemistry to our lives and the environment.

Excel With Objective Chemistry For Iit-Screening Prof. S. K. Khanna, Dr. N. K. Sharma, Dr. B. Kapila, 2006

Physics, Uspekhi, 2007

Competition Science Vision, 1998-08

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Electron Configuration 148 Success Secrets - 148 Most Asked Questions on Electron Configuration - What You Need to Know Dale Rich, 2014-10-15

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