

Honors Chemistry Final Exam Review- 2023

- Make sure you have the following: scientific calculator and a pencil or pen with blue/black ink.
- The final exam will count as a double summative grade.
- The final exam is based on content from 2<sup>nd</sup> semester.

Moles

- How many particles are in a mole?  $6.02 \times 10^{23}$
- What is the conversion between moles and mass?  $\text{moles} = \frac{\text{mass}}{\text{molar mass}}$
- What is the conversion between moles and number of particles?  $\text{moles} = 6.02 \times 10^{23}$
- What is the conversion between moles and volume of gas at STP?  $\text{moles} = 22.4 \text{ L}$
- Do the following conversions:
  - How many grams are in 0.354 mol of cholesterol,  $\text{C}_{27}\text{H}_{46}\text{O}$ ?  $137.9$
  - How many formula units are in 12.3 g potassium nitrate? How many atoms?  $\text{KNO}_3$   $7.32 \times 10^{22}$  formula units
  - What volume at STP would 3.2g of carbon dioxide occupy?  $\text{CO}_2$   $1.6 \text{ L}$
  - How many molecules of helium are in 2.0L balloon at STP?  $\text{He}$   $5.4 \times 10^{22} \text{ molec.}$
  - How many atoms are present in 3.2L of carbon dioxide at STP?  $\text{CO}_2$   $8.6 \times 10^{22} \text{ atoms}$
- Do the following percent composition problems:
  - What is the percent of carbon in carbon dioxide?  $27.29\%$
  - In a sample of pitchblende, uranium ore, it is determined that there are 2.5 g of uranium in a 500.0 g sample. What percent of the ore is uranium?  $0.5\%$
  - What mass of carbon can you expect to recover from a 5.00g sample of glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ )? How many carbon atoms are present?  $12.09\text{g}$   $1.00 \times 10^{23} \text{ atoms}$
  - A piece of jewelry is composed of 0.977mol of metal atoms. If 80.0% of the atoms are silver, how many silver atoms are in the jewelry?  $0.782 \text{ mol}$   $4.71 \times 10^{23} \text{ atoms}$
- Work the following molecular/empirical formula problems:
  - What is the empirical formula for a compound if 4.00g is zinc and 1.00g is oxygen?  $\text{ZnO}$
  - A gaseous compound has a percentage composition of 82.4% nitrogen and 17.6 % hydrogen. What is the empirical formula of the gas?  $\text{NH}_3$
  - What is the molecular formula for a compound with an empirical formula of  $\text{NO}_2$  and a molecular mass of 92.0g/mol?  $\text{N}_2\text{O}_4$
  - What is the empirical formula for a compound if a 32.8g sample contains 22.8g of barium, 2.00g of carbon, and 8.00 g of oxygen?  $\text{BaCO}_3$
  - Identify which of the following are empirical formulas.
 

$\text{H}_2\text{O}$	$\text{CuH}_2$	$\text{H}_2\text{O}_2$	$\text{CH}_2\text{O}$	$\text{Na}_2\text{PO}_4$
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  - A 10.0g sample of sodium carbonate hydrate is heated. After heating, 3.71g of the anhydrous salt remains. What is the formula for this hydrate? Name this hydrate.

$\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$   
Sodium carbonate decahydrate

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