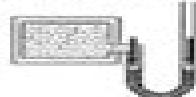


Manometers

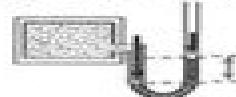
Chem Worksheet

Name _____

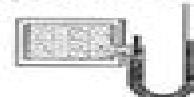
A manometer is a device that measures the pressure of a gas in an enclosed container. It is made from a U-shaped tube filled with mercury. The pressure of the gas in the container is compared to the pressure from the atmosphere. If the gas pressure is the same as the atmospheric pressure, the level of mercury in both sides of the U-tube will be the same. If the gas is at a higher pressure than the atmosphere the mercury level on the side open to the atmosphere will be higher. If the gas is at a lower pressure than the atmosphere the mercury level on the side open to the atmosphere will be lower.



Gas pressure is equal to atmospheric pressure



Gas pressure is more than atmospheric pressure



Gas pressure is less than atmospheric pressure

USEFUL EQUATIONS

$$\begin{aligned}1.00 \text{ atm} &= 101300 \text{ Pa} \\1.00 \text{ atm} &= 760 \text{ torr}\end{aligned}$$

$$\begin{aligned}1.00 \text{ atm} &= 101.3 \text{ kPa} \\1.00 \text{ atm} &= 760 \text{ mmHg}\end{aligned}$$

$$\begin{aligned}1.00 \text{ atm} &= 14.7 \text{ psi} \\1 \text{ cm} &= 10 \text{ mm}\end{aligned}$$

example

An enclosed container of gas is connected to a manometer. The mercury level is 8 cm lower on the side connected to the gas sample. If atmospheric pressure is .984 atm find the pressure of the gas in the container.

- draw a picture

- convert all measurements to mm Hg

$$\frac{8 \text{ cm}}{1 \text{ cm}} \times \frac{10 \text{ mm}}{1 \text{ cm}} = 80 \text{ mm}$$

$$\frac{.984 \text{ atm}}{1 \text{ atm}} \times \frac{760 \text{ mmHg}}{1.00 \text{ atm}} = 748 \text{ mmHg}$$

- add or subtract the measurements based on the drawing

$$748 \text{ mmHg} + 80 \text{ mmHg} = 828 \text{ mmHg}$$



Solve the following problems. Draw a picture of the manometer for each problem.

- What is the pressure of the neon gas sample in the manometer shown to the right?
- A container of helium is connected to a manometer and the mercury level is 145 mm lower on the side open to the atmosphere. Atmospheric pressure is 775 mm Hg. Find the pressure of the helium.
- The mercury in a manometer is 38 mm lower on the side connected to sample of oxygen gas. If the atmospheric pressure is 95.2 kPa determine the pressure of the oxygen.
- What is the pressure of the carbon dioxide in the manometer shown to the right?
- The atmosphere has a pressure of 680 torr. An air-filled container has a pressure of 18.9 PSI and is connected to a manometer. Draw a picture of the manometer and determine the height of the mercury column supported by the air.
- A basketball is attached to a manometer and the mercury is 18 mm higher on the side connected to the atmosphere. The pressure of the atmosphere is 0.95 atm. Find the pressure in the basketball.
- What is the pressure in pascals for the air sample in the manometer pictured to the right?
- A gas container is connected to a manometer. The mercury in the manometer is 7.2 cm lower on the side open to the atmosphere. Atmospheric pressure is measured to be 755 mm Hg. What is the pressure of the gas in atmospheres?



Manometer Practice Problems With Answers

John Bird

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Advanced Biology Michael Roberts, Michael Reiss, Grace Monger, 2000 The major new course text has been written by experienced authors to provide coverage of the Advanced Subsidiary AS and Advanced GCE Biology and Human Biology specifications in a single book Advanced Biology provides clear well illustrated information which will help develop a full understanding of biological structure and function and of relevant applications The topics have been carefully organised into parts which give a logical sequence to the book This new text has been developed to replace the best selling titles Biology Principles and Processes and Biology A Functional Approach Features include full colour design with clear diagrams and photographs up to date information on biotechnology health applied genetics and ecology clearly written text using the latest Institute of Biology terminology a useful summary and a bank of practice questions at the end of every chapter support boxes help bridge the gap from GCSE or equivalent courses extension boxes providing additional depth of content some by guest authors who are experts in their field and a comprehensive index so you can quickly locate information with ease There is also a website providing additional support that you can access directly at www.advancedbiology.co.uk

Prentice Hall Chemistry, 2000 **Fundamental Principles of Veterinary Anesthesia** Gareth E. Zeiler, Daniel S. J. Pang, 2024-01-31

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