

MATH SKILLS TRANSPARENCY

3 Use with
Chapter 3.

Finding Percent by Mass

	Carbon	Hydrogen	Oxygen	Total	Sucrose
Carbon	8.44 g carbon	1.30 g hydrogen	10.26 g oxygen	20.00 g	51.30%
Hydrogen	20.00 g sucrose	20.00 g sucrose	20.00 g sucrose		42.2%
Oxygen					6.50%
Total					100%

Diagram illustrating the percent composition of sucrose (C₁₂H₂₂O₁₁). The total mass of sucrose is 20.00 g. The mass of each element is: Carbon (8.44 g), Hydrogen (1.30 g), and Oxygen (10.26 g). The percent composition is calculated as follows:

- Carbon: $\frac{8.44 \text{ g carbon}}{20.00 \text{ g sucrose}} \times 100\% = 42.2\%$
- Hydrogen: $\frac{1.30 \text{ g hydrogen}}{20.00 \text{ g sucrose}} \times 100\% = 6.50\%$
- Oxygen: $\frac{10.26 \text{ g oxygen}}{20.00 \text{ g sucrose}} \times 100\% = 51.30\%$

Percentages are also shown in a pie chart:

The chart is a pie chart representing the percent composition of sucrose. It is divided into three segments: a large grey segment representing oxygen at 51.30%, a smaller grey segment representing hydrogen at 6.50%, and a medium grey segment representing carbon at 42.2%.

Math Skills Transparency Answers 15

Lauren Gardner

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