

9 • Bonding & Molecular Structure

LEWIS STRUCTURE PRACTICE - 1

Ca^{2+} $[Ca]^{2+}$	NO_3^- $\leftrightarrow \left[\begin{array}{c} \text{O}^- - \text{N} = \text{O} \\ \\ \text{O}^- \end{array} \right]^- \leftrightarrow$	PCl_5
C_2H_2 $\text{H} : \text{C} : : : \text{C} : \text{H}$	S^{2-} $[\text{:}\ddot{\text{S}}\text{:}]^{2-}$	NH_3 $\text{H} : \ddot{\text{N}} : \text{H}$ $ $ H
SO_4^{2-} $\left[\begin{array}{c} \text{O} \\ \vdots \\ \text{O} : \text{S} : \text{O} \\ \vdots \\ \text{O} \end{array} \right]^{2-}$	SO_2 $\text{O}^- - \text{S} = \text{O} \leftrightarrow \text{O} = \text{S} - \text{O}^-$	CaH_2 $\text{H} : \text{Ca} : \text{H}$



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LEWIS STRUCTURE PRACTICE - 2

CH_4 $\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$	HCN $\text{H} : \text{C} : : : \text{N} :$	K^+ $[K]^+$
PO_4^{3-} $\left[\begin{array}{c} \text{O} \\ \vdots \\ \text{O} : \text{P} : \text{O} \\ \vdots \\ \text{O} \end{array} \right]^{3-}$	O_3 $\text{O}^- - \text{O} = \text{O} \leftrightarrow \text{O} = \text{O} - \text{O}^-$	SF_6
Br^- $[\text{:}\ddot{\text{Br}}\text{:}]^-$	AlH_3 $\begin{array}{c} \text{H} : \text{Al} : \text{H} \\ \\ \text{H} \end{array}$	NH_4^+ $\left[\begin{array}{c} \text{H} \\ \\ \text{H} : \text{N} : \text{H} \\ \\ \text{H} \end{array} \right]^+$

Lewis Structure Answer Key

Daniel C. Pantaleo

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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