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web arrhenius bases an arrhenius base is defined as any species that increases the concentration of hydroxide ions oh in aqueous solution an example of an arrhenius base is the highly soluble sodium hydroxide naoh sodium hydroxide dissociates in water as follows $\text{NaOH} \rightleftharpoons \text{Na}^+ + \text{OH}^-$ in water sodium hydroxide

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web work done by svante arrhenius in the late 19th century set the groundwork for our current understanding of acid base theory despite the relatively limited technology available to him he was able to develop theories that adequately explained the properties and chemical reactivity of many acids and bases

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web the arrhenius theory of acids and bases acids are substances that produce hydrogen ions in solution bases are substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to produce water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$ hydrochloric

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web aug 15 2020 the equilibrium formulated in scheme 1 between hydrogen ion and the corresponding acid and base can be called a simple acid base equilibrium by mixing two simple systems a double acid base system and an acid base equilibrium result that can always be formulated as follows $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web key points a brønsted lowry acid is any species that is capable of donating a proton H^+ a brønsted lowry base is any species that is capable of accepting a proton which requires a lone pair of electrons to bond to the H^+ water is amphoteric which means it can act as both a brønsted lowry acid and a brønsted lowry base

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web the A^- ion is a base because it is accepting a proton from the H_2O the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid HA loses a proton it forms a base A^- when the base A^- accepts a proton back again it obviously reforms the acid HA

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web his theory gave a generalized explanation of acids and bases based on structure and bonding through the use of the lewis definition of acids and bases chemists are now able to predict a wider variety of acid base reactions lewis theory used electrons instead of proton transfer and specifically stated that an acid is a species that accepts

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