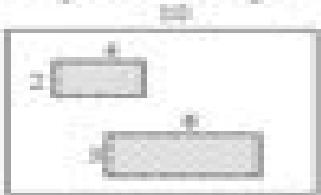


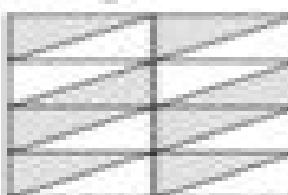
## Geometric Probability -- Please graph these on your own graph paper.

Find the probability of hitting the shaded area or length of the target with a randomly thrown dart that hits the target. All measurements are in centimeters. The formula is **Geometric Probability =  $\frac{\text{shaded area or shaded length}}{\text{total area or total length}}$** . For full credit, you MUST SHOW the formulas and calculations of the shaded area or length, the total area or length, and then reduced fractions.

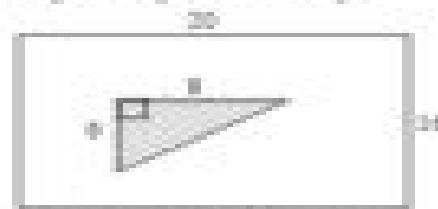
1. rectangle in a rectangle



2. rectangle in a square



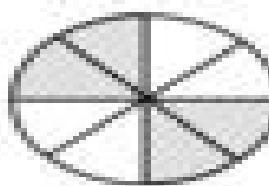
3. right triangle in a rectangle



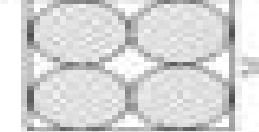
4. circle in a square



5. circle divided into 8 equal pieces



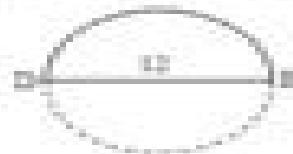
6. three circles inside a square



7. If a point falls on the perimeter, what is the probability of it landing on the hypotenuse of the right triangle?



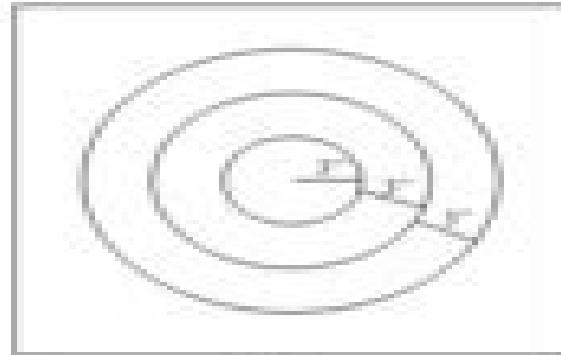
8. If a point falls on the outer edge of the top half of the circle (positioned by a solid line below), what is the probability it is on the curve?



9. On  $\overline{AC}$ , what is the probability of hitting a blue bag at random and landing within 1 cm of the point B?



10.



If a randomly thrown dart lands within the square (not centered), what is the probability:

a. that it will land within the middle circle?

b. that it will land outside the largest circle?

c. that it will land between the largest and middle circles?

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**Paul J. Campbell, Louise S. Grinstein**

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web what is the mole fraction of  $\text{LiCl}$  in the solution 11.8 practice problem acetone  $\text{C}_3\text{H}_6\text{O}$  has a density of 0.791 g/ml and xylene  $\text{C}_8\text{H}_{10}$  has a density of 0.864 g/ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19.0 ml of acetone was dissolved in 98.4 ml of xylene 7.1

chemteam gas law dalton s law problems 1 10 - Feb 24 2022

web since the mole fraction of the  $\text{O}_2$  is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21. 0.050/4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature