



MTY1103 ANALYTICAL CHEMISTRY

Name _____ Section _____ Score _____

CHAPTER 6 WORKSHEET

- Balance the following reactions in both acidic and basic media
 - $Hg_2^{2+} + H_2 \rightleftharpoons Hg + H^+$
 - $Cr_2O_7^{2-} + I^- \rightarrow Cr^{3+} + I_2$
 - $MnO_4^- + C_2O_4^{2-} \rightarrow MnO_2 + CO_3^{2-}$
- Solve for the cell potential of the following cells
 - $Zn|Zn^{2+}||Cu^{2+}|Cu$
 - $Hg_2^{2+} + H_2 \rightleftharpoons Hg + H^+$
 - $3Fe^{2+} + 2Al \rightarrow 3Fe + 2Al^{3+}$
 - $Fe^{2+} + Cr_2O_7^{2-} \rightarrow Fe^{3+} + Cr^{3+}$
 - $Fe^{2+} + MnO_4^- \rightarrow Fe^{3+} + Mn^{2+}$
- Solve for the equilibrium constant and the free energy change of the following reactions
 - $Sn + 2Cu^{2+} \rightleftharpoons Sn^{2+} + 2Cu^+$
 - $Fe^{2+} + 2Ag \rightleftharpoons Fe + 2Ag^+$
 - $Zn + Cu^{2+} \rightleftharpoons Zn^{2+} + Cu$
 - $Au|Au(NO_3)_3||Co(NO_3)_2|Co$
 - $Sn^{2+}, Sn^{4+}||Hg^{2+}, Hg_2^{2+}$
 - $Zn|Zn^{2+}||Ag^+|Ag$
- Calculate the E_{cell} of the reaction $Co(s) + Fe^{2+}(aq) \rightarrow Co^{2+}(aq) + Fe(s)$ given that $[Co^{2+}] = 0.25 M$ and $[Fe^{2+}] = 0.94 M$.
- Calculate the E_{cell} of the reaction $Cd(s) + Fe^{2+}(aq) \rightarrow Cd^{2+}(aq) + Fe(s)$ given that $[Fe^{2+}] = 0.60 M$ and $[Cd^{2+}] = 0.010 M$.
- Consider the galvanic cell shown below. In a certain experiment, the emf of the cell is found to be 0.54 V at 25°C. Suppose that $[Zn^{2+}] = 2.5 M$ and $P_{H_2} = 1.0 atm$, what is the pH of the solution?

$$Zn(s) + 2H^+(? M) \rightarrow Zn^{2+}(2.5 M) + H_2(1.0 atm)$$
- What is the emf of a galvanic cell consisting of a Cd^{2+}/Cd half-cell and a $Pt/H^+/H_2$ half-cell if $[Cd^{2+}] = 0.20 M$, $[H^+] = 0.16 M$, and $P_{H_2} = 0.80 atm$?

Episode 1103 Chemistry Worksheet

O García



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