

**Please remember that simply knowing or memorizing the information below is not sufficient. You must be able to know when and how to APPLY the items below appropriately.*

5.1 Basics

Pressure and Units (*see guide!*)

5.2 Simple Gas Laws (*see guide!*)

5.3 Ideal Gas Law (*see preview below!*)

Used to determine one gas variable for a given scenario, knowing the other variables.

A. $PV=nRT$ ($R = 0.0821 \text{ L atm/mol K}$; $8.314 \text{ L kPa/mol K}$; be sure to match all units to the constant)

B. Limitations:

1. Works well at low pressures *as the molecules of the gas rarely collide or come into close enough proximity for intermolecular forces to be significant*
2. Works well at high temperatures *as the molecules of the gas have so much energy that the intermolecular forces and energy lost in collisions is negligible*
3. Most gases do not behave ideally above 1 atm pressure
4. Does not work well near the condensation conditions of a gas
5. Deviations occur due to stronger attractive forces

C. Derivation to relate molar mass and density

$D = m/V$ $n = m/M$ where M is molar mass

$PV = m/MRT$; $M = DRT/P$

5.4 Gas Mixtures and Collecting Gas of Water

5.5 Diffusion and Effusion

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