

## Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

1)	nitrogen	$1s^2 2s^2 2p^3$
2)	copper	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$
3)	chlorine	$1s^2 2s^2 2p^6 3s^2 3p^5$
4)	calcium	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
5)	lead	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

In the space below, write the noble gas notation electron configurations of the following elements:

6)	barium	$[Xe] 6s^2$
7)	gold	$[Xe] 4f^{14} 5d^{10} 6s^1$
8)	platinum	$[Xe] 4f^{14} 5d^8 6s^1$
9)	xenon	$[Kr] 4d^{10} 5s^2 5p^6$
10)	calcium	$[Ar] 4s^2$

Determine what elements are denoted by the following electron configurations:

11)	$1s^2 2s^2 2p^6 3s^2 3p^3$	phosphorous
12)	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$	ruthenium
13)	$[Kr] 4d^{10} 5s^2 5p^6$	iodine
14)	$[Xe] 4f^{14} 5d^{10} 6s^2$	mercury
15)	$[Rn] 7s^2$	radium

Determine which of the following electron configurations are not correct:

16)	$1s^2 2s^2 2p^6 3s^2 3p^5 4s^2$	3p <sup>5</sup> should be a 6
17)	$1s^2 2s^2 2p^6 3s^2 2p^5$	1s only holds 2 and there aren't two "2p" orbitals
18)	$[Cs] 5f^{14} 6d^1 7s^2$	Cs is not a noble gas
19)	$[Kr] 4d^{10} 5s^2 5p^4$	Correct
20)	$[Ne]$ for Neon	Can't use a noble gas to shortcut itself

# Gpb Chemistry Electron Distribution Answers

**Robert D. Prim**

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