

## Skills Worksheet

# Concept Review

## Section: Calculating Quantities in Reactions

Complete each statement below by writing the correct term or phrase.

1. All stoichiometric calculations involving equations use \_\_\_\_\_ ratios.
2. When solving stoichiometric problems, you must \_\_\_\_\_ the equation first.
3. Balanced equations give the \_\_\_\_\_ numbers of moles of substances.
4. \_\_\_\_\_ in chemical equations provide mole ratios that can be used as conversion factors.
5. The conversion factor for converting between mass and moles is the \_\_\_\_\_ of the substance.
6. In making calculations involving \_\_\_\_\_, you must convert volume to mass.
7. To convert from volume to mass, you can use the \_\_\_\_\_ of the substance as the conversion factor.
8. When calculating the number of particles, you can use \_\_\_\_\_ as the conversion factor.

In the blanks at the left, write the letter of the choice that best completes the statement or answers the question. Consider the following problem when answering:

What mass of sulfuric acid is required to neutralize 2.65 g of potassium hydroxide? The products of the reaction are potassium sulfate and water.

- \_\_\_\_\_ 9. What should you do first after reading the problem carefully?
  - a. Estimate the answer.
  - b. Calculate the molar mass of sulfuric acid.
  - c. Write a balanced chemical equation.
  - d. Convert all masses to moles.
- \_\_\_\_\_ 10. What should you do before setting up the problem?
  - a. Determine the densities.
  - b. Calculate molar masses.
  - c. Convert all masses to moles.
  - d. Estimate the answer.

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