

Equilibrium Reaction Problems (2019-2020)

1.1 The equilibrium constant K_c for the reaction below is 0.0040 at 250°C.



For each of the following, write an expression for the equilibrium constant. The reactants are on the left, the products are on the right.

(a) $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ (b) $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ (c) $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$

(d) $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$ (e) $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ (f) $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$

(g) $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$ (h) $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ (i) $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$

1.2 At 250°C, K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 0.0040. Calculate the value of K_c for the reaction $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$ at 250°C.



1.3 At 250°C, K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 0.0040. Calculate the value of K_c for the reaction $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ at 250°C.



1.4 At 250°C, K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 0.0040. Calculate the value of K_c for the reaction $\text{NO}_2(g) \rightleftharpoons \frac{1}{2}\text{N}_2\text{O}_4(g)$ at 250°C.



1.5 At 250°C, K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 0.0040. Calculate the value of K_c for the reaction $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ at 250°C.



1.6 At 250°C, K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 0.0040.



Calculate the equilibrium constant K_c for the reaction $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ at 250°C. (Note: The products are on the right, the reactants are on the left.)

(a) $K_c = 0.0040$ (b) $K_c = 0.0020$ (c) $K_c = 0.0010$ (d) $K_c = 0.0005$

(e) $K_c = 0.00025$ (f) $K_c = 0.000125$ (g) $K_c = 0.0000625$

1.7 At 250°C, K_c for the reaction $2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ is 0.0040.



Calculate the equilibrium constant K_c for the reaction $\text{N}_2\text{O}_4(g) \rightleftharpoons 2\text{NO}_2(g)$ at 250°C. (Note: The products are on the right, the reactants are on the left.)

(a) $K_c = 0.0040$ (b) $K_c = 0.0020$ (c) $K_c = 0.0010$ (d) $K_c = 0.0005$

(e) $K_c = 0.00025$ (f) $K_c = 0.000125$ (g) $K_c = 0.0000625$

Equilibrium Shift Practice Problems

Michael, Robin Bade, Michael Parkin



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