

 Honors Chemistry: Chapter 4 - Problem Set (with some 6)

- All answers and work on a separate sheet of paper!

Classify the following as always true (AT), sometimes true (ST), or never true (NT)

1. Atoms of one element change into atoms of another element during chemical reactions.
2. Atoms combine in one-to-one ratios to form compounds.
3. Atoms of one element are different from atoms of other elements.
4. According to Dalton's atomic theory, atoms are composed of protons, electrons and neutrons.
5. Atoms of elements are electrically neutral.
6. The mass of an electron is equal to the mass of a neutron.
7. The charge on all protons is the same.
8. The atomic number of an element is the sum of the protons and electrons in an atom of that element.
9. An atom of nitrogen has 7 protons and 7 neutrons.
10. Relative atomic masses are expressed in amu's.
11. The number of neutrons in the nucleus can be calculated by subtracting the atomic number from the mass number.
12. In his periodic table, Mendeleev arranged the element in ascending order of atomic number.
13. The representative elements are the Group A elements.
14. The transition metals and inner transition metals are the Group B elements.
15. The element in Group 4A, period 3 is gallium.

Match each description in Column B with the correct term in Column A.

Column A

1. atom
2. Dalton's atomic theory
3. Democritius
4. electrons
5. protons
6. neutrons
7. nucleus
8. atomic number
9. mass number
10. isotopes
11. atomic mass unit (amu)
12. atomic mass
13. periodic table
14. periods
15. group
16. representative elements
17. alkali metals
18. transition metals
19. halogens
20. noble gases

Column B

- a. the group 7A elements
- b. horizontal rows of the periodic table
- c. Group A elements
- d. arrangement of elements according to similarities in their properties
- e. a portion of the B Group elements
- f. Group 1A elements
- g. Group 0 elements or group 8A
- h. vertical column on the periodic table
- i. represents the number of protons in an atom
- j. 1/12 of the mass of the carbon -12 atom
- k. total number of protons and neutrons in an atom
- l. weighted average mass of the atoms in a naturally occurring sample of an element
- m. atoms that have the same number of protons but different numbers of neutrons
- n. positively charged subatomic particles
- o. subatomic particles with no charge
- p. negatively charged subatomic particles
- q. the central positively charged core of the atom
- r. teacher in ancient Greece
- s. theory to explain the reactive behavior of elements
- t. smallest particle of an element that retains the properties of that element

Solve the following problems:

Living By Chemistry Answer Key Unit 4

RJ Shavelson

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expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems
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