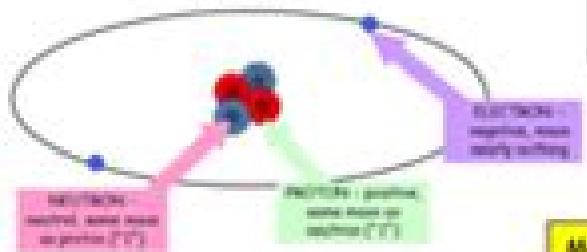


The Atom

Year 11 C2 - Discovering Chemistry

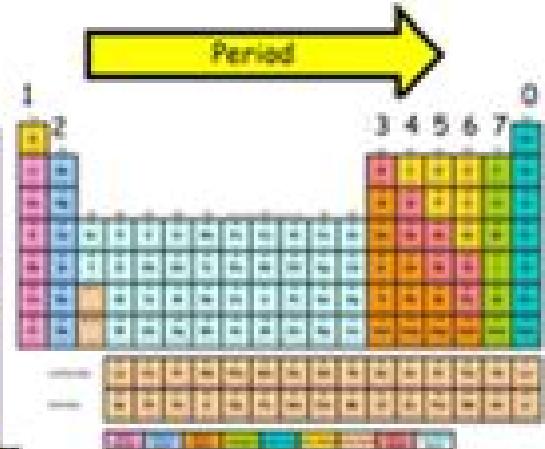
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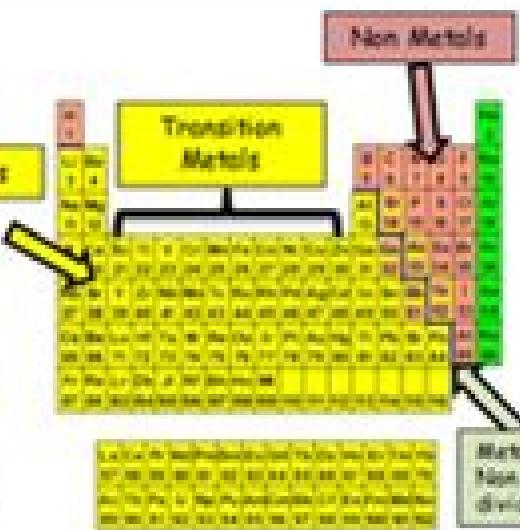
Atomic Number	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-1

He

$$\begin{array}{c} 4 \\ \leftarrow \text{Mass number} = \text{number of protons} + \text{number of neutrons} \\ \text{Orbits} \\ \leftarrow \text{Mass number} = \text{number of protons} \end{array}$$



(Most important) All of the elements in the same group have similar **PROPERTIES**. That is how I thought of the periodic table in the first place. This is called **PERIODICITY**.



How can an atom have a **decimal** for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 as the average atomic mass is:

$$35 \times 0.75 + 37 \times 0.25 = 35.5$$

I came up with the first periodic table. I arranged the elements in order of increasing atomic mass.



I put elements with similar properties into horizontal rows. I also left gaps because any new elements were discovered.

N
7

Mg
12

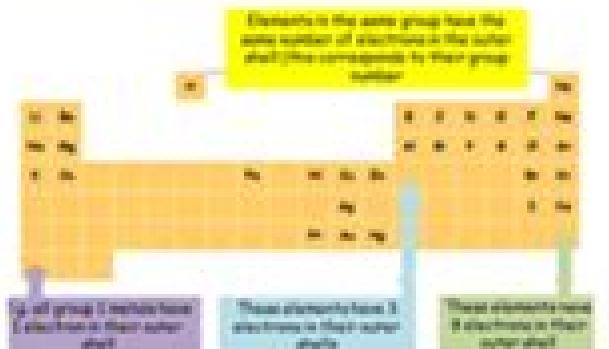
Ca
20



Electron structure
+ 1,2

Electron structure
+ 2,8,2

Electron structure
+ 2,8,8,2



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