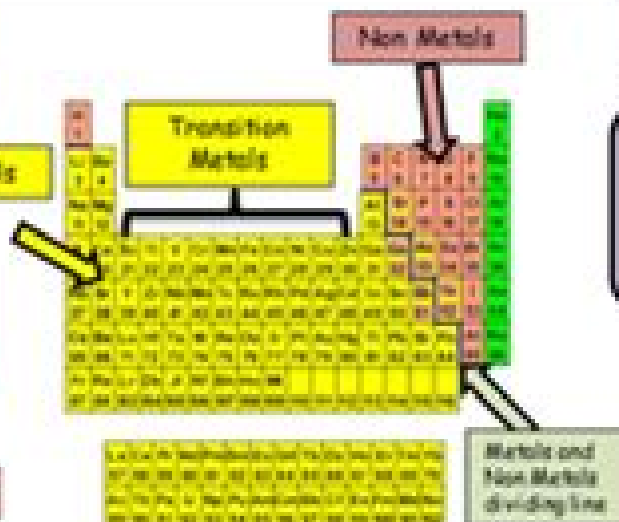


Year 11 C2 - Discovering Chemistry Topic 1



Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-



I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

Dmitri Mendeleev



I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

Isotopes

Higher

Example is an atom with a different number of neutrons

Notice that the mass number is different. How many neutrons does each isotope have?

16 17 18

Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more

N 14 / 7

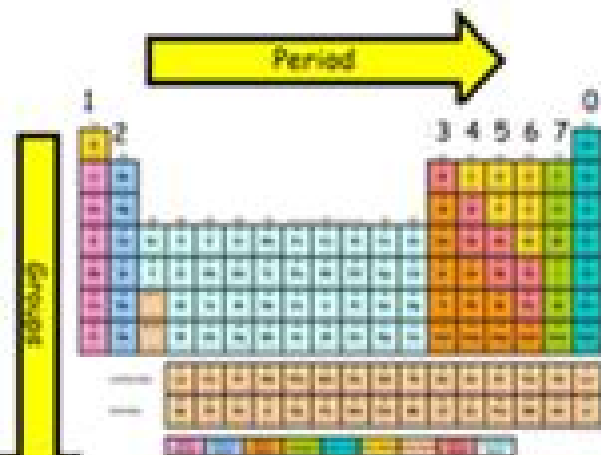
Electron structure + 2,5

Mg 24 / 12

Electron structure + 2,8,2

Ca 40 / 20

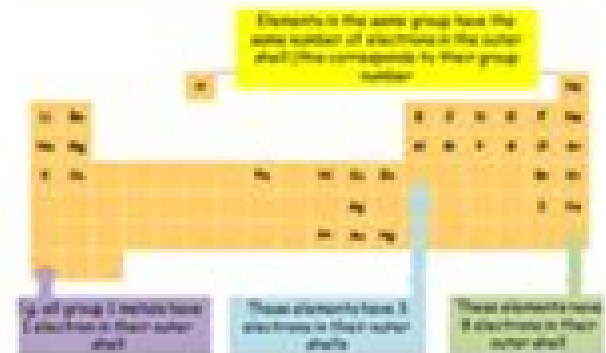
Electron structure + 2,8,8,2



How can an atom have a **decimal** for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$35 \times 3 + 1 \times 37 = 106$



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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G.C.S.E.



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 - The Rise of Digital Reading Gcse Edexcel Chemistry C2 2013
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Edexcel Chemistry C2 2013
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Edexcel Chemistry C2 2013
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Edexcel Chemistry C2 2013
 - Personalized Recommendations
 - Gcse Edexcel Chemistry C2 2013 User Reviews and Ratings
 - Gcse Edexcel Chemistry C2 2013 and Bestseller Lists
5. Accessing Gcse Edexcel Chemistry C2 2013 Free and Paid eBooks
 - Gcse Edexcel Chemistry C2 2013 Public Domain eBooks
 - Gcse Edexcel Chemistry C2 2013 eBook Subscription Services
 - Gcse Edexcel Chemistry C2 2013 Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Gcse Edexcel Chemistry C2 2013 Compatibility with Devices
 - Gcse Edexcel Chemistry C2 2013 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gcse Edexcel Chemistry C2 2013
 - Highlighting and Note-Taking Gcse Edexcel Chemistry C2 2013
 - Interactive Elements Gcse Edexcel Chemistry C2 2013
8. Staying Engaged with Gcse Edexcel Chemistry C2 2013
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Edexcel Chemistry C2 2013
9. Balancing eBooks and Physical Books Gcse Edexcel Chemistry C2 2013
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Edexcel Chemistry C2 2013
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gcse Edexcel Chemistry C2 2013
 - Setting Reading Goals Gcse Edexcel Chemistry C2 2013
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Edexcel Chemistry C2 2013
 - Fact-Checking eBook Content of Gcse Edexcel Chemistry C2 2013
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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