

Report Sheet

Use the table of standard reduction potentials from your textbook to complete the following table. The zinc-copper Daniell cell is given as an example.

Electrodes	Conc. of Each Cation	Half-Reactions	Cathode or Anode?	E° (table)	E°_{cell} (calc'd)	E°_{cell} (measured)
A. Determine the E° for the Standard Zn-Cu Galvanic Cell						
Zn	1.0 M Zn^{2+}	$Zn^{2+}_{\text{aq}} + 2e^- \rightarrow Zn_{\text{ss}}$	Anode	-0.76 V	+1.30 V	0.93 V
Cu	1.0 M Cu^{2+}	$Cu^{2+}_{\text{aq}} + 2e^- \rightarrow Cu_{\text{ss}}$	Cathode	+0.34 V		
B. Determine the E° for the Cu-Zn Cell at Non-standard Concentrations						
Zn	1.0 M Zn^{2+}	$Zn^{2+}_{\text{aq}} + 2e^- \rightarrow Zn_{\text{ss}}$	Anode	-0.76 V		0.73 V
Cu	0.5 M Cu^{2+}	$Cu^{2+}_{\text{aq}} + 2e^- \rightarrow Cu_{\text{ss}}$	Cathode	+0.34 V		
Zn	1.0 M Zn^{2+}	$Zn^{2+}_{\text{aq}} + 2e^- \rightarrow Zn_{\text{ss}}$	Anode	-0.76 V		0.96 V
Cu	0.1 M Cu^{2+}	$Cu^{2+}_{\text{aq}} + 2e^- \rightarrow Cu_{\text{ss}}$	Cathode	+0.34 V		
C. Determine the E° for Zn-Mg and Cu-Mg Galvanic Cells						
Zn	1.0 M Zn^{2+}	$Zn^{2+}_{\text{aq}} + 2e^- \rightarrow Zn_{\text{ss}}$	Anode	-0.76 V		0.82 V
Mg	1.0 M Mg^{2+}	$Mg^{2+}_{\text{aq}} + 2e^- \rightarrow Mg_{\text{ss}}$	Cathode	-2.38 V		
Cu	1.0 M Cu^{2+}	$Cu^{2+}_{\text{aq}} + 2e^- \rightarrow Cu_{\text{ss}}$	Anode	-0.76 V		1.10 V
Mg	1.0 M Mg^{2+}	$Mg^{2+}_{\text{aq}} + 2e^- \rightarrow Mg_{\text{ss}}$	Cathode	-2.38 V		

*You must use the Nernst equation to calculate this E_{cell} voltage. (See Equation 8.)

Galvanic Cell Experiment Report

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