

Name: _____ Date: _____ Class: _____

1. What kind of paper is coated with the blue litmus paper shown below?

A. a base
B. an acid
C. water
D. neutral



2. The pH of the liquid in contact with the blue litmus paper in the picture is _____.

A. 7
B. 14
C. 9
D. 6



3. The fact that vinegar has a pH and is acidic only it tastes _____.

A. sour
B. salty
C. sour
D. bitter



4. The salt used to flavor food is made up of _____.

A. two metals
B. two non-metals
C. a metal and a non-metal
D. hydrogen and a non-metal



5. Using the pH scale shown below, determine the pH of 200 mL rainwater.



A. between 7 and 14
B. between 1 and 7
C. between 12 and 14
D. 7

6. Blood plasma usually has a pH between _____.

A. 1 and 3
B. 4 and 6
C. 11 and 14
D. 7 and 9

pH Scale



7. What is the reactivity order in decreasing pH on the pH scale?



A. 1 and 2 hydroxide and 3 base
B. 4 is a good acid
C. 5 has equal parts of H and OH
D. 6 is not a good acid

8. What is the salt formed when the elements of Na₂O and H₂O are mixed together?

A. sodium oxide
B. sodium hydroxide
C. sodium nitrate
D. sodium chloride



Acids, Bases & Salts

Acids & Their Properties

An acid is a pure substance that dissociates into hydrogen ions (H⁺) and an anion. Acids are typically strong electrolytes that dissociate completely in aqueous solutions. Acids have a pH less than 7.



Hydrochloric acid dissociates into hydrogen ions and chloride ions when it is dissolved in water.



Water dissociates into hydrogen ions and hydroxide ions when it is dissolved in water.

Characteristics of Acids

- Have a sour taste
- Turn blue litmus red
- React with bases to form salt and water
- React with metals to form salt and hydrogen gas



Acids turn blue litmus red.

Acids react with bases to form salt and water.

Acids react with metals to form salt and hydrogen gas.

Acids are strong electrolytes that dissociate into hydrogen ions when they are dissolved in water.

Acids have a pH less than 7.

Acids taste sour.

Acids are strong acids.

Bases & Their Properties

A base is a pure substance that dissociates into hydroxide ions (OH⁻) and a metal cation. Bases are typically strong electrolytes that dissociate completely in aqueous solutions.



Hydrochloric acid dissociates into sodium ions and hydroxide ions when it is dissolved in water.



Water dissociates into hydrogen ions and hydroxide ions when it is dissolved in water.

Strength of Acids & Bases

Acids and bases range in strength decreasing as the number of hydroxide ions (OH⁻) increases. Strong bases have many hydroxide ions (OH⁻) and therefore have a high pH. Weak bases have few hydroxide ions (OH⁻) and therefore have a low pH.



The pH Scale

Locate the pH scale to measure the strength of acids and bases. The pH scale ranges from 1 to 14. Bases with pH less than 7 are called acids. The lower the pH, the stronger the acid. Bases with pH greater than 7 are called bases. The higher the pH, the stronger the base.



Characteristics of Bases

- Have a bitter taste
- Indicators
- Change litmus paper to blue



Salts

Acid-Base Reactions

An acid-base reaction occurs when an acid and base react to form a salt and water. The hydroxide ions (OH⁻) from one reactant react with the hydrogen ions (H⁺) from the other reactant to form water.



reactants products



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Table of Contents Guide Acids Bases And Salts

1. Understanding the eBook Guide Acids Bases And Salts
 - The Rise of Digital Reading Guide Acids Bases And Salts
 - Advantages of eBooks Over Traditional Books
2. Identifying Guide Acids Bases And Salts
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guide Acids Bases And Salts
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guide Acids Bases And Salts
 - Personalized Recommendations
 - Guide Acids Bases And Salts User Reviews and Ratings
 - Guide Acids Bases And Salts and Bestseller Lists

5. Accessing Guide Acids Bases And Salts Free and Paid eBooks
 - Guide Acids Bases And Salts Public Domain eBooks
 - Guide Acids Bases And Salts eBook Subscription Services
 - Guide Acids Bases And Salts Budget-Friendly Options
6. Navigating Guide Acids Bases And Salts eBook Formats
 - ePUB, PDF, MOBI, and More
 - Guide Acids Bases And Salts Compatibility with Devices
 - Guide Acids Bases And Salts Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guide Acids Bases And Salts
 - Highlighting and Note-Taking Guide Acids Bases And Salts
 - Interactive Elements Guide Acids Bases And Salts
8. Staying Engaged with Guide Acids Bases And Salts
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guide Acids Bases And Salts
9. Balancing eBooks and Physical Books Guide Acids Bases And Salts
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guide Acids Bases And Salts
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guide Acids Bases And Salts
 - Setting Reading Goals Guide Acids Bases And Salts
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guide Acids Bases And Salts
 - Fact-Checking eBook Content of Guide Acids Bases And Salts
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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