

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Class: \_\_\_\_\_

1. What kind of liquid is in contact with the **blue litmus paper** (litmus paper)?

- A. a base
- B. salt
- C. water
- D. an acid



2. The **pH** of the liquid in contact with the **blue litmus paper** in the picture is \_\_\_\_\_.

- A. 7
- B. 14
- C. 0
- D. 0



3. The fact that vinegar has positive acid in it suggests why it tastes \_\_\_\_\_.

- A. sweet
- B. salty
- C. sour
- D. bitter



4. The **salt** used to flavor food is made up of \_\_\_\_\_.

- A. two metals
- B. two nonmetals
- C. a metal and a nonmetal
- D. hydrogen and a nonmetal



5. Using the **pH scale** shown below, determine the pH of **acid rain**.



- A. between 7 and 14
- B. between 1 and 7
- C. between 12 and 14
- D. 7

6. **Alkaline** bases usually have a pH between \_\_\_\_\_.

- A. 1 and 3
- B. 4 and 6
- C. 12 and 14
- D. 7 and 9



7. What is the reason why water is considered **neutral** on the pH scale?



- A. it is the both an acid and a base
- B. it is a good solvent
- C. it has equal parts of H and OH
- D. it is not a good solvent

8. What is the **salt** formed when the cations of **limes** (or **lithium**) are a **metal cation**?

- A. sodium oxide
- B. sodium hydroxide
- C. sodium iodide
- D. sodium chloride



## Acids & Their Properties

An **acid** is a substance that releases **hydrogen ions** ( $H^+$ ) when dissolved in water. Acids are **sour** and **corrosive**. They can **react** with **metals** to produce **hydrogen gas**. Acids are **found** in many **foods** and **drinks**. They are also **found** in many **household products**.



Hydrogen ions and chloride ions are released when hydrochloric acid is dissolved in water.



Hydrogen ion ( $H^+$ )

## Characteristics of Acids

- Taste sour (lemons)
- Turn blue litmus red
- React with metals to produce hydrogen gas
- Conduct electricity

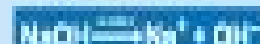


Acids are **found** in many **foods** and **drinks**. They are also **found** in many **household products**.



## Bases & Their Properties

A **base** is a substance that releases **hydroxide ions** ( $OH^-$ ) when dissolved in water. Bases are **bitter** and **corrosive**. They can **react** with **acids** to produce **water** and **salts**.



The **sodium** ion and **hydroxide** ion are released when sodium hydroxide is dissolved in water.



Hydroxide ion ( $OH^-$ )

## Characteristics of Bases

- Taste bitter (soap)
- Turn red litmus blue
- Conduct electricity



## Strength of Acids & Bases

Acids and bases are **strong** or **weak** depending on the number of **hydrogen ions** ( $H^+$ ) or **hydroxide ions** ( $OH^-$ ) they release. **Strong** acids and bases release a **large** number of ions, while **weak** acids and bases release a **small** number of ions.



## The pH Scale

The **pH scale** is a measure of the **strength** of acids and bases. The scale ranges from 1 to 14. A solution with a pH of 7 is **neutral**. A solution with a pH of 1 is **strongly acidic**. A solution with a pH of 14 is **strongly basic**.

## Salts

### Acid-Base Reactions

An **acid-base reaction** occurs when an acid and a base react to form a **salt** and **water**. The **hydrogen ions** ( $H^+$ ) from the acid combine with the **hydroxide ions** ( $OH^-$ ) from the base to form **water**.



# Guide Acids Bases And Salts

**John McMurry, Susan McMurry**



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1. Understanding the eBook Guide Acids Bases And Salts
  - The Rise of Digital Reading Guide Acids Bases And Salts
  - Advantages of eBooks Over Traditional Books
2. Identifying Guide Acids Bases And Salts
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guide Acids Bases And Salts
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guide Acids Bases And Salts
  - Personalized Recommendations
  - Guide Acids Bases And Salts User Reviews and Ratings
  - Guide Acids Bases And Salts and Bestseller Lists

5. Accessing Guide Acids Bases And Salts Free and Paid eBooks
  - Guide Acids Bases And Salts Public Domain eBooks
  - Guide Acids Bases And Salts eBook Subscription Services
  - Guide Acids Bases And Salts Budget-Friendly Options
6. Navigating Guide Acids Bases And Salts eBook Formats
  - ePub, PDF, MOBI, and More
  - Guide Acids Bases And Salts Compatibility with Devices
  - Guide Acids Bases And Salts Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guide Acids Bases And Salts
  - Highlighting and Note-Taking Guide Acids Bases And Salts
  - Interactive Elements Guide Acids Bases And Salts
8. Staying Engaged with Guide Acids Bases And Salts
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guide Acids Bases And Salts
9. Balancing eBooks and Physical Books Guide Acids Bases And Salts
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guide Acids Bases And Salts
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guide Acids Bases And Salts
  - Setting Reading Goals Guide Acids Bases And Salts
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guide Acids Bases And Salts
  - Fact-Checking eBook Content of Guide Acids Bases And Salts
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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