

BOARD QUESTION PAPER: MARCH 2022

CHEMISTRY

Times 2 Min

Max. Marks: 70

General Instructions:

The question paper is divided into four sections.

(1) **Section A:** Q. No. 1 contains Ten multiple choice type of questions carrying One mark each.
Q. No. 2 contains Eight very short answer type of questions carrying One mark each.

(2) **Section B:** Q. No. 3 to Q. No. 14 are Twelve short answer type of questions carrying Two marks each. (Attempt any Eight).

(3) **Section C:** Q. No. 15 to Q. No. 26 are Twelve short answer type of questions carrying Three marks each. (Attempt any Eight).

(4) **Section D:** Q. No. 27 to Q. No. 31 are Five long answer type of questions carrying Four marks each. (Attempt any Three).

(5) Use of log table is allowed. Use of calculator is not allowed.

(6) Figures to the right indicate full marks.

(7) For each multiple choice type of question, it is mandatory to write the correct answer along with its alphabet, e.g. (a) ... (b) ... (c) ... (d) ... etc.

*No marks(s) shall be given, if ONLY the correct answer or the alphabet of the correct answer is written.
Only the first attempt will be considered for evaluation.*

SECTION - A

Q.1. Select and write the correct answer for the following multiple choice type of questions: [10]

- The co-ordination number of atoms in body centred cubic structure (bcc) is _____.
(a) 4 (b) 6 (c) 8 (d) 12
- In calculating osmotic pressure, the concentration of solute is expressed in _____.
(a) molarity (b) molality
(c) mole fraction (d) percentage mass
- The enthalpy change for the chemical reaction $\text{H}_2\text{O}_{(l)} \longrightarrow \text{H}_2\text{O}_{(g)}$ is called enthalpy of
(a) vapourisation (b) fusion
(c) combustion (d) sublimation
- Which of the following transition element shows maximum oxidation state?
(a) Sc (b) Fe (c) Mn (d) V
- The correct formula for the complex compound, sodium hexacyanoferrate (III) is _____.
(a) $\text{Na}[\text{Fe}(\text{CN})_6]$ (b) $\text{Na}_2[\text{Fe}(\text{CN})_6]$
(c) $\text{Na}_3[\text{Fe}(\text{CN})_6]$ (d) $\text{Na}_4[\text{Fe}(\text{CN})_6]$

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