

## Interpreting the Periodic Table

Examine the hypothetical periodic table below and answer the given questions.

A 10x10 grid with various shaded regions and labels. The grid is composed of 100 squares. Shaded regions include a 2x3 rectangle in the top-left, a 3x2 rectangle in the top-right, a 2x2 square in the center, a 2x2 square in the bottom-right, and a 3x3 square in the bottom-right corner. Labels are placed in specific squares: 'A' is in the bottom-left square (row 10, column 1), 'B' is in the bottom-right square (row 10, column 10), 'C' is in the center square (row 5, column 5), 'D' is in the top-center square (row 2, column 5), 'E' is in the top-right square (row 1, column 10), 'F' is in the middle-right square (row 7, column 10), 'G' is in the bottom-middle square (row 9, column 5), 'H' is in the middle-left square (row 7, column 1), 'I' is in the top-middle square (row 2, column 2), 'K' is in the top-left square (row 1, column 2), 'L' is in the middle-right square (row 7, column 2), 'P' is in the bottom-right square (row 10, column 2), and 'Z' is in the bottom-right corner square (row 10, column 10).

① Which pair of elements has the same number of valence electrons?  
\_\_\_\_\_

② Which pair of metals are in the same period? \_\_\_\_\_

③ Which element has the smallest atomic number? \_\_\_\_\_

④ Which elements are metals? \_\_\_\_\_

⑤ Which elements are nonmetals? \_\_\_\_\_

⑥ If the atomic number of the element P is 41, what is the atomic number of the element P?

# Exam Questions Introduction To Periodic Table

**Neale Taylor, Angela Stubbs, Robert Stokes**

## Exam Questions Introduction To Periodic Table:

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