

Lab: Single and Double Replacement Reactions

Purpose: To predict products and observe various single and double replacement reactions.

Hypotheses: Single Replacement Reactions: Use the **activity series** (p.288 – Horner) and your understanding of single replacement reactions to predict products for the following reactions. If no reaction will occur, then write "no rxn" on the product side of the equation. Remember to **BALANCE** the equation afterwards.

- 1) $\text{HCl} + \text{Mg} \rightarrow$ _____
- 2) $\text{HCl} + \text{Cu} \rightarrow$ _____ (Hint: copper could be +2 in this compound)
- 3) $\text{HCl} + \text{Fe} \rightarrow$ _____ (Hint: iron will be +3 in this compound)
- 4) $\text{HCl} + \text{Zn} \rightarrow$ _____

Hypotheses: Double Replacement Reactions: Use the **solubility chart** and your understanding of double replacement reactions to predict products for the following reactions (see Table 2 on the back of this paper...we're only fully writing out equations for 4 of the 16 you will test). If both of the possible products are aqueous, then write "no rxn" on the product side of the equation. Otherwise, clearly state which product will be aqueous and which will form a solid precipitate. Remember to **BALANCE** the equation afterwards.

- 1) $\text{Na}_2\text{CO}_3(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow$ _____
- 6) $\text{NaOH}(\text{aq}) + \text{Fe}(\text{NO}_3)_3(\text{aq}) \rightarrow$ _____
- 11) $\text{Na}_3\text{PO}_4(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$ _____
- 16) $\text{MgSO}_4(\text{aq}) + \text{Zn}(\text{NO}_3)_2(\text{aq}) \rightarrow$ _____

Procedures:

1. For each single replacement reaction, place a sample of each metal in your well plate and then place 5-6 drops of HCl (aq) on top of it. Record your observations.
2. For each double replacement reaction, place two drops of each solution carefully on the transparency, covering each "X" appropriately over your "Super High-Tech Patented Reaction Surface." Record your observations.
3. If a precipitate formed, clearly write the formula for the precipitate on the line provided.
4. Clean up your workstation according to your instructors directions and answer all questions.

Observations/Data:

Table 1: Single Replacement Reaction Observations

	HCl (aq)
Mg	1 Observations:
Cu	2 Observations:
Fe	3 Observations:
Zn	4 Observations:

Lab 27 Double Replacement Reactions

James Signorelli

Lab 27 Double Replacement Reactions:

CHEMISTRY EXPERIMENTS James Signorelli,2014-09-19 Gifted and talented students and any student interested in pursuing a science major in college needs a rigorous program to prepare them while they are still in high school This book utilizes a format where the application of several disciplines science math and language arts principles are mandated Each lab concludes with either an essay or a detailed analysis of what happened and why it happened This format is based on the expectations of joining a university program or becoming an industrial science professional The ideal student lab report would be written in a lab research notebook and then the essay or final analysis is done on a word processor to allow for repeat editing and corrections The research notebook has all graph pages a title section and a place for the students and their assistants to sign and witness that exercise The basic mechanics of the lab report title purpose procedure diagrams data table math and calculations observations and graphs are handwritten into the book The conclusion is done on a word processor MS Word which allows the instructor to guide the student in writing and editing a complete essay using the MLA format When the final copy is completed the essay is printed and inserted into the lab notebook for grading At the end of the term the student has all their labs in one place for future reference These lab notebooks can be obtained for as little as 3 00 per book This is money well spent In our district the Board of Education buys the books for each student The BOE sees these books as expendable but necessary materials for all science and engineering instruction

Chemistry Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02 **Top Shelf** Brian Pressley,2003 Covers chemical formulas and equations chemical reactions structure of atoms the gas laws and more Presents hands on activities as catalysts to fuel student imagination

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