

## Lab – Specific Heat of a Metal

The **specific heat capacity** of a substance is the amount of heat energy necessary to raise the temperature of 1 gram of a substance by 1°C. It may be thought of as “heat storing ability.” Every substance has its own specific heat capacity – this physical property relates energy change to temperature change. Materials with low specific heat values need low energy to increase/decrease temperature. The same is true of materials with high specific heat – they need high energy to increase/decrease temperature.

Metals have relatively low specific heats. On the other hand, the specific heat of water is relatively high: **1 cal / g • °C**. Water will absorb much energy before its temperature increases even slightly. This helps to explain why the climate near a large lake is generally warmer in the winter and cooler in the summer, and why your body (which is nearly 60% water) can tolerate rapid changes in external temperatures.

The relationship between heat and temperature is shown in Big Equation #1 of Thermochemistry:

**Energy = (mass of material) x (its change in temp) x (its specific heat)**

$$q = m \times C \times \Delta T$$

Finally, remember that calorimeters work because they let no heat in or out. As a result, the heat energy gained by the water in a calorimeter is assumed to be equal to the heat loss by the metal:

$$\Delta H_{\text{water}} = -\Delta H_{\text{metal}}$$

### PROCEDURE:

1. Select a metal sample and record its mass.
2. Create a “Hot Water Bath” using a 400 mL beaker. Fill the beaker with tap water so that the metal will be totally submerged in the bath. Bring the water to a boil.
3. Use the loop end of a test tube brush to carefully transfer the metal sample into the boiling water and **continue to boil for an additional 5 minutes** (to ensure that the piece of metal reaches the same temperature as the boiling water).
4. In the meantime, assemble your calorimeter. Place a Styrofoam cup into a 400 mL beaker. Have a thermometer and cardboard lid ready. Finally, measure **exactly** 100 mL of tap water in a graduated cylinder and carefully pour this water into the calorimeter. *Remember, due to water’s density, 1 mL of water = 1 g of water.*
5. Record the temperature of the cool tap water in the calorimeter. This is the initial temperature of water.
6. Following the 5-minute boiling period of your metal sample, record the temperature of the boiling water. This is the initial temperature of the metal.
7. Turn off Bunsen burner to stop heating the water.
8. Carefully transfer the piece of heated metal to the calorimeter and cap with lid. Place the thermometer through the center of the lid and be sure the tip is submerged in the water. **DO NOT PUSH TOO HARD OR THE CALORIMETER MAY BREAK!**
9. Gently swirl the sample and container. Watch the thermometer and record the highest temperature reached. This is the final temperature for both the water and the metal.
10. Answer the discussion questions. Note that you will be working in calories. 1 cal = 4.184 J **OR** 1 J = 0.239 cal

### DATA TABLE:

WATER DATA		METAL DATA	
Mass of tap water		Mass of metal	
Initial temperature of water (°C)		Initial temperature of metal (°C)	
Final temperature of water (°C)		Final temperature of metal (°C)	
Specific Heat Capacity of Water (cal/g•°C)		Specific Heat Capacity of Metal (cal/g•°C)	<i>TO BE CALCULATED (in Question # 3)</i>

# Lab 22 Specific Heat Of Metals Answers

**Christian G. Meyer**

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