



### Magnitude of the Equilibrium Constant

For a reaction

$$aA + bB \rightleftharpoons cC + dD$$

the equilibrium constant  $K$  is defined as the ratio of products to reactants at equilibrium:

$$K = \frac{c^c d^d}{a^a b^b}$$

Because  $K$  represents the ratio of products to reactants, the magnitude of  $K$  is an indicator of the extent of products and reactants present when the reaction is at equilibrium. Some chemical reactions proceed almost fully to products, whereas other chemical reactions begin with reactants at  $100\%$  and tend to have significant amounts of both reactants and products at equilibrium.

#### Part A

Hydrogen is used as a reducing agent because it is very light and reacts inexpensively and conveniently with oxygen. For the combustion reaction  $2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$ , what is the likely magnitude of the equilibrium constant  $K$ ?

• None

• Hint 1: How to approach the problem: [View Hint](#)

•  $K < 10^{-4}$

•  $10^{-4} \leq K \leq 10^0$

•  $K = 1$

•  $K > 10^2$

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It is clear that  $100\%$  product must be formed in the reaction.

#### Part B

The industrial production of iron (Fe) from hematite (Fe<sub>2</sub>O<sub>3</sub>) is represented in the following reaction:



Temperature (K)	0
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**Kenneth L. Higbee**

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