

## Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

### Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

### Model 1 – Ion Charges for Selected Elements

1	H <sup>+</sup>											
2	Li <sup>+</sup>	Be <sup>2+</sup>						N <sup>3-</sup>	O <sup>2-</sup>	F <sup>1-</sup>		
3	Na <sup>+</sup>	Mg <sup>2+</sup>						P <sup>3-</sup>	S <sup>2-</sup>	Cl <sup>1-</sup>		
Transition elements												
4	K <sup>+</sup>	Ca <sup>2+</sup>	Fe <sup>2+</sup> Fe <sup>3+</sup>	Ni <sup>2+</sup> Ni <sup>3+</sup>	Cu <sup>+</sup> Cu <sup>2+</sup>	Zn <sup>2+</sup>					Br <sup>1-</sup>	
5	Rb <sup>+</sup>	Sr <sup>2+</sup>			Ag <sup>1+</sup>			Se <sup>2-</sup> Se <sup>4-</sup>		I <sup>-</sup>		
6		Ba <sup>2+</sup>				Hg <sub>2</sub> <sup>2-</sup> Hg <sup>2+</sup>		Pb <sup>2+</sup> Pb <sup>4+</sup>				
Cations →						← Anions						

1. Based on the information in Model 1:

a. Identify three elements that form only one cation. *→ Positive Ions*  
 $\text{Na}^+$ ,  $\text{Sr}^{2+}$ ,  $\text{Al}^{3+}$  (Any box with one element works.)

b. Identify three elements that form only one anion. *→ Negative Ions*  
 $\text{P}^{3-}$ ,  $\text{O}^{2-}$ ,  $\text{Br}^{1-}$  (Any box with one element works.)

c. Identify three elements that form more than one cation.

$\text{Fe}^{+2}$  and  $\text{Fe}^{+3}$ ,  $\text{Ni}^{+2}$  and  $\text{Ni}^{+3}$ ,  $\text{Ca}^{+1}$  and  $\text{Ca}^{+2}$

d. In what region of the periodic table are these "multiple ion" elements usually located?

*Transition elements or d block*

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.

$2\text{K}^+$  and  $5^{2-}$  or  $\text{K}_2\text{S}^{-2}$

3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.

$\text{Fe}^{2+}$  and  $5^{2-}$  ( $\text{FeS}$ ) and  $2\text{Fe}^{3+}$  and  $3\text{S}^{2-}$  ( $\text{Fe}_2\text{S}_3$ )



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## **Table of Contents Labpaq Chemistry Answer Key**

1. Understanding the eBook Labpaq Chemistry Answer Key
  - The Rise of Digital Reading Labpaq Chemistry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Labpaq Chemistry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Labpaq Chemistry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Labpaq Chemistry Answer Key
  - Personalized Recommendations

- Labpaq Chemistry Answer Key User Reviews and Ratings
- Labpaq Chemistry Answer Key and Bestseller Lists

5. Accessing Labpaq Chemistry Answer Key Free and Paid eBooks

- Labpaq Chemistry Answer Key Public Domain eBooks
- Labpaq Chemistry Answer Key eBook Subscription Services
- Labpaq Chemistry Answer Key Budget-Friendly Options

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- ePUB, PDF, MOBI, and More
- Labpaq Chemistry Answer Key Compatibility with Devices
- Labpaq Chemistry Answer Key Enhanced eBook Features

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- Adjustable Fonts and Text Sizes of Labpaq Chemistry Answer Key
- Highlighting and Note-Taking Labpaq Chemistry Answer Key
- Interactive Elements Labpaq Chemistry Answer Key

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- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Labpaq Chemistry Answer Key

9. Balancing eBooks and Physical Books Labpaq Chemistry Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Labpaq Chemistry Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Labpaq Chemistry Answer Key

- Setting Reading Goals Labpaq Chemistry Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Labpaq Chemistry Answer Key

- Fact-Checking eBook Content of Labpaq Chemistry Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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