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$$4\text{NO(g)} + 6\text{H}_2\text{O(l)} + 4\text{NH}_3\text{(g)} + 5\text{O}_2\text{(g)} \rightarrow 11\text{H}_2\text{O(l)} + 7\text{N}_2\text{(g)}$$
 a is this reaction endothermic or exothermic
 b how much heat will be absorbed or released if 10 g of NO is reacted with 10 g of H_2O

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web heats of fusion and vaporization determines what determines how much energy is required to change a substance's phase
 what is the heat of fusion formula
 $q = n \cdot \Delta H_{\text{fus}}$
 what is the heat of vaporization formula
 $q = n \cdot \Delta H_{\text{vap}}$
 what is on the x axis of the heating curve
 time

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 $2\text{C(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{CO}_2\text{(g)}$
 is this endothermic or exothermic

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 $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(l)}$
 energy endothermic the transfer of heat from the surroundings to the system
 $2\text{Hg(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{HgO(s)}$

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web q 140 g of 184 J/mol OC_2H_5 9372 J/mol $\text{C}_2\text{H}_5\text{OH}$ 1 mole KCN 0.5453 moles 65.1 g of 0.5453 mole moles 5 when a 15 g sample of SrCl_2 dissolves in 80 g of water in a calorimeter the temperature drops from 42°C to 15°C calculate the ΔH for the process
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