

## Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- 1)  $\text{C}_6\text{H}_6$  **CH**
- 2)  $\text{C}_6\text{H}_{14}$   **$\text{C}_6\text{H}_{12}$**
- 3)  $\text{WO}_3$   **$\text{WO}_3$**
- 4)  $\text{C}_2\text{H}_4\text{O}_2$   **$\text{CH}_2\text{O}$**
- 5)  $\text{X}_2\text{Y}_{12}$   **$\text{X}_2\text{Y}$**
- 6) A compound with an empirical formula of  $\text{C}_2\text{OH}_4$  and a molar mass of 88 grams per mole. What is the molecular formula of this compound?  **$\text{C}_2\text{O}_2\text{H}_4$**
- 7) A compound with an empirical formula of  $\text{C}_2\text{H}_4\text{O}$  and a molar mass of 136 grams per mole. What is the molecular formula of this compound?  **$\text{C}_2\text{H}_4\text{O}_2$**
- 8) A compound with an empirical formula of  $\text{CFBrO}$  and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound?  **$\text{C}_2\text{F}_2\text{Br}_2\text{O}_2$**
- 9) A compound with an empirical formula of  $\text{C}_2\text{H}_4\text{N}$  and a molar mass of 46 grams per mole. What is the molecular formula of this compound?  **$\text{C}_2\text{H}_4\text{N}$**
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula  $\text{C}_2\text{H}_4\text{NO}$ . If its molar mass is 116.1 g/mol, what is the molecular formula of the compound?  **$\text{C}_2\text{H}_4\text{N}_2\text{O}_2$**
- 11) A certain blue solid contains 36.84% N and 63.16% O. What is the empirical formula of this compound? The ratios are  $\text{N}_{1.50}\text{O}_{1.50}$ . Since 1.50 is not close to a whole number, we multiply both subscripts by 2. The empirical formula is thus  $\text{N}_2\text{O}_3$ . (The name is dinitrogen trioxide.)  **$\text{N}_2\text{O}_3$**
- 12) A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound?  **$\text{InCl}_3$**
- 13) An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula?  **$\text{K}_2\text{C}_2\text{O}_4$**
- 14) Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol?  **$\text{C}_3\text{H}_8\text{O}$**

# Empirical Formula Answer Key

**Disha Experts**

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This book delves into Empirical Formula Answer Key. Empirical Formula Answer Key is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Empirical Formula Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Empirical Formula Answer Key
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- Chapter 3: Empirical Formula Answer Key in Everyday Life
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This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Empirical Formula Answer Key.

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## Table of Contents Empirical Formula Answer Key

1. Understanding the eBook Empirical Formula Answer Key
  - The Rise of Digital Reading Empirical Formula Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Empirical Formula Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Answer Key
  - Personalized Recommendations
  - Empirical Formula Answer Key User Reviews and Ratings
  - Empirical Formula Answer Key and Bestseller Lists
5. Accessing Empirical Formula Answer Key Free and Paid eBooks
  - Empirical Formula Answer Key Public Domain eBooks
  - Empirical Formula Answer Key eBook Subscription Services
  - Empirical Formula Answer Key Budget-Friendly Options
6. Navigating Empirical Formula Answer Key eBook Formats
  - ePUB, PDF, MOBI, and More
  - Empirical Formula Answer Key Compatibility with Devices
  - Empirical Formula Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Empirical Formula Answer Key
  - Highlighting and Note-Taking Empirical Formula Answer Key
  - Interactive Elements Empirical Formula Answer Key
8. Staying Engaged with Empirical Formula Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Empirical Formula Answer Key

9. Balancing eBooks and Physical Books Empirical Formula Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Empirical Formula Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Empirical Formula Answer Key

- Setting Reading Goals Empirical Formula Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Empirical Formula Answer Key

- Fact-Checking eBook Content of Empirical Formula Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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