

Substance	Lewis Structure	Molar Mass	Intermolecular Attractive Force
<b>Methanol</b> $\text{CH}_3\text{OH}$ $12 + 3(1) + 16 + 1 = 32$	$\begin{array}{c} \text{H} \\   \\ \text{H} - \text{C} - \ddot{\text{O}} - \text{H} \\   \\ \text{H} \end{array}$	32.042	Hydrogen Bonds
<b>Ethanol</b> $\text{CH}_3\text{CH}_2\text{OH}$ $12 + 5 + 16 + 1 + 1 = 30$	$\begin{array}{c} \text{H} \quad \text{H} \quad   \\   \quad   \quad   \\ \text{H} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\   \quad   \\ \text{H} \quad \text{H} \end{array}$	46.068	Hydrogen Bonds
<b>n-Butanol</b> $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\   \quad   \quad   \quad   \quad   \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\   \quad   \quad   \quad   \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array}$	74.122	Hydrogen Bonds
<b>Glycerin</b> $\text{C}_3\text{H}_8(\text{OH})_3$ $12 + 8 + 18 + 3 = 39$	$\begin{array}{c} \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \\   \quad   \quad   \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{H} \\   \quad   \quad   \\ \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \end{array}$	92.093	Hydrogen Bonds
<b>Water</b> $\text{H}_2\text{O}$	$\begin{array}{c} \text{H} \\   \\ \text{H} - \ddot{\text{O}} - \text{H} \\   \\ \text{H} \end{array}$	18.015	Hydrogen Bonds

# Exploring Intermolecular Forces Lab Answers

**Lauren Gardner**



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