

Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

(Note to teachers and students: This Gizmo was designed as a follow-up to the Bohr Model: Introduction Gizmo. We recommend doing that activity before trying this one.)

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- What happens to an electron when it absorbs a tiny packet of light called a **photon**?
the electron absorbs the energy from it, causing a higher energy level
- What happens to an electron when it emits a photon?
the electrons lose energy, causing a lower energy level

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.

The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the **Bohr Model of Hydrogen** Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

- With the **Laser energy** set to 7.0 eV, click **Play** (▶). Observe the blue electron on the **ORBITALS** pane. Do 7.0-eV photons have any effect on the electron? **No, it doesn't**, the electron stays at **lvl 1**.
- Click **Pause** (⏸). Set the **Laser energy** to 12.1 eV and click **Play**. What happens to the electron now? **the electron shakes up and down and moves up by 2 orbital (level 3)**
- Under **Go to energy level**, select 1. Notice that a photon is emitted by the electron. What is the energy of the emitted photon? **(12.1 eV)**

Explore Learning Exploration Guide Bohr Model

Answers

Bruce Cameron Reed

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The Bohr Atom Bruce Cameron Reed,2020-11-11 All students of physics encounter the Bohr model of the atom However it is often covered quickly in order that curricula can progress to wave mechanics This book gives students and instructors a fuller exploration to Bohr s model Topics covered include the historical background to the model Bohr s approach to his original derivation and corollary issues

such as the role of angular momentum in the theory ionized helium the correspondence principle the fine structure constant de Broglie matter waves application of the theory to the diatomic hydrogen molecule and the magnetic field created by the orbiting electron It also includes student exercises a bibliography a list of important physical constants and a survey of Bohr s subsequent life and career Key Features Provides a full historical background to Bohr s model a detailed derivation and corollary calculations associated with the model Serves as a supplementary text to undergraduate physics courses Contains exercises and a list of important physical constants

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Table of Contents Explore Learning Exploration Guide Bohr Model Answers

1. Understanding the eBook Explore Learning Exploration Guide Bohr Model Answers
 - The Rise of Digital Reading Explore Learning Exploration Guide Bohr Model Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Exploration Guide Bohr Model Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Exploration Guide Bohr Model Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Exploration Guide Bohr Model Answers

- Personalized Recommendations
- Explore Learning Exploration Guide Bohr Model Answers User Reviews and Ratings
- Explore Learning Exploration Guide Bohr Model Answers and Bestseller Lists

5. Accessing Explore Learning Exploration Guide Bohr Model Answers Free and Paid eBooks

- Explore Learning Exploration Guide Bohr Model Answers Public Domain eBooks
- Explore Learning Exploration Guide Bohr Model Answers eBook Subscription Services
- Explore Learning Exploration Guide Bohr Model Answers Budget-Friendly Options

6. Navigating Explore Learning Exploration Guide Bohr Model Answers eBook Formats

- ePub, PDF, MOBI, and More
- Explore Learning Exploration Guide Bohr Model Answers Compatibility with Devices
- Explore Learning Exploration Guide Bohr Model Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Explore Learning Exploration Guide Bohr Model Answers
- Highlighting and Note-Taking Explore Learning Exploration Guide Bohr Model Answers
- Interactive Elements Explore Learning Exploration Guide Bohr Model Answers

8. Staying Engaged with Explore Learning Exploration Guide Bohr Model Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Explore Learning Exploration Guide Bohr Model Answers

9. Balancing eBooks and Physical Books Explore Learning Exploration Guide Bohr Model Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Explore Learning Exploration Guide Bohr Model Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Explore Learning Exploration Guide Bohr Model Answers

- Setting Reading Goals Explore Learning Exploration Guide Bohr Model Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Explore Learning Exploration Guide Bohr Model Answers

- Fact-Checking eBook Content of Explore Learning Exploration Guide Bohr Model Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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