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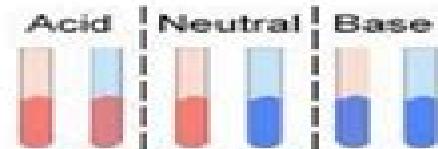
Student Exploration: Titration

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: acid, analyte, base, dissociate, equivalence point, indicator, litmus paper, molarity, neutralize, pH, strong acid, strong base, titrant, titration, titration curve, weak acid, weak base

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

There are several definitions of acids and bases. According to the Brønsted-Lowry definition, an **acid** is a substance that is capable of donating a proton to another substance. A **base** is a substance that accepts protons. When an acid and a base are combined, the acid is **neutralized** as the base accepts the protons produced by the acid.



One way to determine if a solution is acidic or basic is to use **litmus paper**, as shown above. There are two types of litmus papers: red and blue.

How does litmus paper indicate an acid?

Both strips turn red.

How does litmus paper indicate a neutral substance?

The red strip stays red and the blue one stays blue.

How does litmus paper indicate a base?

Both strips turn blue.

Gizmo Warm-up

Litmus is an example of an **indicator**, a substance that changes color depending on its **pH** (pH is a measure of the concentration of protons, or H^+ ions). In the Titration Gizmo, you will use indicators to show how acids are neutralized by bases, and vice versa.

To begin, check that **1.00 M NaOH** is selected for the **Burette**, **Mystery HBr** is selected for the **Flask**, and **Bromthymol blue** is selected for the **Indicator**.

1. Look at the flask. What is the color of the bromthymol blue indicator?
2. What does this tell you about the pH of the solution in the flask?

yellow

pH is below 6.0



Solutions with a pH below 7.0 are acidic, while those with a pH above 7.0 are basic.

3. Move the slider on the burette to the top to add about 25 mL of NaOH to the flask. What happens, and what does this tell you about the pH of the flask?

The solution turns blue meaning the pH is now above 7.6.

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Table of Contents Explore Learning Gizmo Answers Titration

1. Understanding the eBook Explore Learning Gizmo Answers Titration
 - The Rise of Digital Reading Explore Learning Gizmo Answers Titration
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Gizmo Answers Titration
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Gizmo Answers Titration
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Gizmo Answers Titration
 - Personalized Recommendations

- Explore Learning Gizmo Answers Titration User Reviews and Ratings
- Explore Learning Gizmo Answers Titration and Bestseller Lists

5. Accessing Explore Learning Gizmo Answers Titration Free and Paid eBooks

- Explore Learning Gizmo Answers Titration Public Domain eBooks
- Explore Learning Gizmo Answers Titration eBook Subscription Services
- Explore Learning Gizmo Answers Titration Budget-Friendly Options

6. Navigating Explore Learning Gizmo Answers Titration eBook Formats

- ePUB, PDF, MOBI, and More
- Explore Learning Gizmo Answers Titration Compatibility with Devices
- Explore Learning Gizmo Answers Titration Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Explore Learning Gizmo Answers Titration
- Highlighting and Note-Taking Explore Learning Gizmo Answers Titration
- Interactive Elements Explore Learning Gizmo Answers Titration

8. Staying Engaged with Explore Learning Gizmo Answers Titration

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Explore Learning Gizmo Answers Titration

9. Balancing eBooks and Physical Books Explore Learning Gizmo Answers Titration

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Explore Learning Gizmo Answers Titration

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Explore Learning Gizmo Answers Titration

- Setting Reading Goals Explore Learning Gizmo Answers Titration
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Explore Learning Gizmo Answers Titration

- Fact-Checking eBook Content of Explore Learning Gizmo Answers Titration

- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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