

1. In a chemical reaction atoms combine in different ways to form new products.
2. Whatever the change, you cannot lose any atoms. This is known as the law of conservation of mass. It means that if you could weigh the reactants at the start, and then compare this with the mass of all the products formed at the end, they would be equal.
- 3a. When a metal is heated in air it gains mass in forming the metal oxide. **9Hwf4ans**
- 3b. When a carbonate dissolves in acid, the mixture loses mass.
- 3c. Both these situations involve the 'adding' or 'losing' of a gas, which is rather difficult to weigh!
4. From the 'law' in 2. you can make predictions about the quantities of chemicals involved in a reaction.
- 4a. For example, 24g of magnesium burns in air and combines with 16g of oxygen to form 40g of magnesium oxide, even if you can't weigh the oxygen!
- 4b. This means 36g of magnesium will combine with 24g of oxygen from the air to form 60g of magnesium oxide.
- 4c. If 6g of magnesium is burned it combines with 4g of oxygen to form 10g of magnesium oxide.
- 4d. Burning 12g of magnesium will combine with 8g of oxygen to make 20g of magnesium oxide
- 4e. burning 48g of magnesium will combine with 32g of oxygen to make 80g of magnesium oxide.

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**Keith Johnson, Lawrie Ryan, Sue
Adamson**



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