

Limiting Reactant and Percent Yield Worksheet

Key

1. Calcium hydroxide, used to neutralize acid spills, reacts with hydrochloric acid according to the following unbalanced equation:



- a. If you have spilled 6.3 mol of HCl and put 2.8 mol of Ca(OH)_2 on it, which substance is the limiting reactant?

$$\frac{6.3 \text{ mol HCl}}{2 \text{ mol HCl}} \bigg| \frac{1 \text{ mol CaCl}_2}{2 \text{ mol HCl}} = 3.2 \text{ mol CaCl}_2$$

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{1 \text{ mol CaCl}_2}{1 \text{ mol Ca(OH)}_2} = 2.8 \text{ mol CaCl}_2$$

* Ca(OH)_2 is limiting

- b. How many moles of the excess reactant remain?

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{2 \text{ mol HCl}}{1 \text{ mol Ca(OH)}_2} = 5.6 \text{ mol HCl used}$$

$$6.3 \text{ mol HCl} - 5.6 \text{ mol HCl} = \boxed{0.7 \text{ mol HCl remain}}$$

2. Aluminum oxidizes according to the following unbalanced equation: $\underline{4}\text{Al} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{Al}_2\text{O}_3$

- a. Powdered Al (0.048 mol) is placed into a container containing 0.030 mol O_2 . What is limiting reactant?

$$\frac{0.048 \text{ mol Al}}{4 \text{ mol Al}} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} = 0.024 \text{ mol Al}_2\text{O}_3$$

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} = 0.020 \text{ mol Al}_2\text{O}_3$$

* O_2 is limiting

- b. How many moles of the excess reactant remain?

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = 0.040 \text{ mol Al used}$$

$$0.048 \text{ mol Al} - 0.040 \text{ mol Al} = \boxed{0.008 \text{ mol Al remain}}$$

3. Heating zinc sulfide in the presence of oxygen yields the following: $\underline{2}\text{ZnS} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{ZnO} + \underline{2}\text{SO}_2$

- a. If 1.72 mol of ZnS is heated in the presence of ^{45.6}~~32.0~~ g of O_2 , which is the limiting reactant? (Balance the equation first.)

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{2 \text{ mol SO}_2}{1 \text{ mol ZnS}} = 3.44 \text{ mol SO}_2$$

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} = 0.950 \text{ mol SO}_2$$

* O_2 is limiting

- b. How many grams of the excess reactant remain?

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol ZnS}}{3 \text{ mol O}_2} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 92.6 \text{ g used}$$

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 168 \text{ g ZnS (started with)}$$

$$\frac{168 \text{ g ZnS}}{92.6 \text{ g ZnS}} = \boxed{75 \text{ g ZnS remain}}$$

Limiting Reagent And Percent Yield Answer Key 12 3

Fred M. Dewey



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Foundations of College Chemistry, Alternate Morris Hein, Susan Arena, 2010-01-26 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years this book has helped them master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis *Foundations of College Chemistry* Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis Introduction to General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career *Basic Concepts of Chemistry* Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *Chemistry* John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian

institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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CHEMISTRY SILBERBERG,2003 Exploring Chemistry (Loose-Leaf) Matthew Johll,2012-05-04 Matthew Johll s Exploring Chemistry overs the standard topics for the nonmajors course in the typical order but each chapter unfolds in the context of a single case study that helps students connect what they are learning to real life situations For example students work through the often difficult topics of molecular structure gas laws and organic chemistry by learning about the development of powerful new chemotherapy drugs new technologies for screening airline passengers and the creation of

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mains signal cos 2 pi 60 t 60hz mains frequency

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it is a combination of three deflections q r and s seen on a typical ecg signal where p is the first deflection q is the first
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