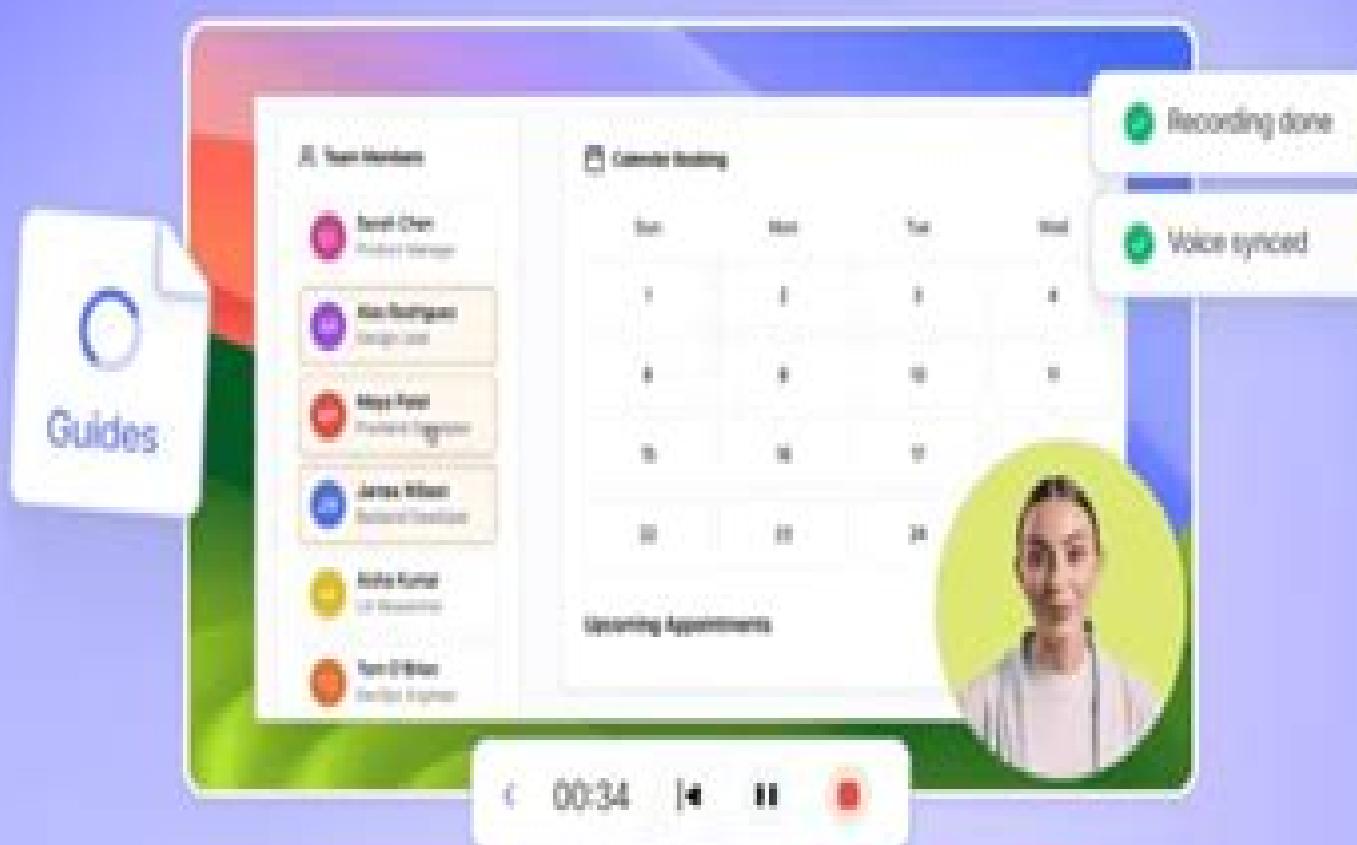


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web the reactant which is in a lesser amount than is required by stoichiometry is the limiting reactant in an alternate method of finding the limiting reagent the amount of product formed by each reactant is calculated the limiting reactant is the reactant from which the minimum amount of product is formed also if we calculate the amount of

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web the limiting reactant or limiting reagent is the reactant that gets consumed first in a chemical reaction and therefore limits how much product can be formed as we saw in example 1 there are many different ways to determine the limiting reactant but they all involve using mole ratios from the balanced chemical equation

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web may 1 2013 the limiting reactant or limiting reagent is the reactant that determines the amount of product that can be formed in a chemical reaction the reaction proceeds until the limiting reactant is completely used up in our example above

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web the key to recognizing which reactant is the limiting reactant is to determine the reactant that yields the least amount of product is the limiting reactant it doesn t matter which product is used for the calculation as long as the same product is