

ENGAGE

Why does the atomic number but not the mass number change in positron emission?

SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission

TRY IT FIRST

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

SOLUTION GUIDE

ANALYZE THE PROBLEM	Given	Need	Connect
	Mn-49, positron emission	balanced nuclear equation	mass number, atomic number of new nucleus

STEP 1 Write the incomplete nuclear equation.



STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = ? + 0$$

$$49 - 0 = ?$$

$49 - 0 = 49$ (mass number of new nucleus)

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = ? + 1$$

$$25 - 1 = ?$$

$25 - 1 = 24$ (atomic number of new nucleus)

STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written $^{49}_{24}\text{Cr}$.

STEP 5 Complete the nuclear equation.



STUDY CHECK 5.4

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

ANSWER



TEST

Try Practice Problems 5.15 to 5.18



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Community and Junior College Journal ,1974

Study Guide Joan I. Senyk,Darrell D. Ebbing,1984

Chemistry Theodore Lawrence Brown,Harold Eugene LeMay,Bruce Edward Bursten,Catherine Jones Murphy,Patrick M. Woodward,Matthew Stoltzfus,2023 Publishing the fifteenth edition of this text bespeaks an exceptionally long record of successful textbook writing We are appreciative of the loyalty and support the book has received over the years and we are mindful of our obligation to justify each new edition We begin our approach to each new edition with an intensive author retreat in which we ask ourselves the deep questions that we must answer before we can move forward What justifies yet another edition What is changing in the world not only of chemistry but with respect to science education and the qualities of the students we serve How can we help your students not only learn the principles of chemistry but also become critical thinkers who can think more like chemists The answers lie only partly in the changing face of chemistry itself The introduction of many new technologies has changed the landscape in the teaching of sciences at all levels The use of online resources in accessing information and presenting learning materials has markedly changed the role of the textbook as one element among many tools for student learning Our challenge as authors is to maintain the text as the primary source of chemical knowledge and practice while at the same time integrating it with the new avenues for learning made possible by technology This book incorporates a number of technologies to improve pedagogy including use of computer based classroom tools such as Mastering Chemistry which is continually evolving to provide more effective means of testing and evaluating student performance while giving the student immediate and helpful feedback Video feedback for a select number of Exam

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