

(ix) The standard potential of the cell in the following reaction is \_\_\_\_\_.  
 $\text{Cd}_{(\text{ss})} + \text{Cu}_{(\text{ss})}^{2+} \longrightarrow \text{Cd}_{(\text{ss})}^{2+} + \text{Cu}_{(\text{ss})}$   
 $(E_{\text{cell}}^{\circ} = -0.403 \text{ V}, E_{\text{cell}}^{\circ} = 0.334 \text{ V})$

(a)  $-0.737 \text{ V}$  (b)  $0.737 \text{ V}$   
 (c)  $-0.069 \text{ V}$  (d)  $0.069 \text{ V}$

(x) The value of  $[\text{H}_3\text{O}^+]$  in mol  $\text{lit}^{-1}$  of 0.001 M acetic acid solution ( $K_a = 1.8 \times 10^{-5}$ ) is \_\_\_\_\_.  
 (a)  $1.34 \times 10^{-1}$  (b)  $1.34 \times 10^{-2}$   
 (c)  $1.34 \times 10^{-3}$  (d)  $1.34 \times 10^{-4}$

**Q.2. Answer the following questions:**

[8]

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:  
 $\text{NO}_{(\text{g})} + \text{O}_{(\text{g})} \longrightarrow \text{NO}_{(\text{g})} + \text{O}_{(\text{g})}$

(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:  
 $[\text{Cu}(\text{NH}_3)_4] [\text{PtCl}_4]$  and  $[\text{Pt}(\text{NH}_3)_4] [\text{CuCl}_4]$

(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for  $\text{CH}_3\text{Br}_{(\text{g})}$  is  $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$  at  $25^\circ\text{C}$ . What is solubility of  $\text{CH}_3\text{Br}_{(\text{g})}$  in water at same temperature and partial pressure of 0.164 bar?

**SECTION – B**

**Attempt any EIGHT of the following questions:**

[16]

**Q.3.** Explain pseudo-first order reaction with suitable example.

**Q.4.** Write the consequences of Schottky defect with reasons.

**Q.5.** What is the action of following on ethyl bromide:  
 (i) Na in dry ether  
 (ii) Mg in dry ether

**Q.6.** Explain formation of peptide linkage in protein with an example.

**Q.7.** Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

**Q.8.** Explain monodentate and ambidentate ligands with example.

**Q.9.** Explain the trends in the following atomic properties of group 16 elements:

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