

## Introduction to Acids & Bases Webquest

Please complete this webquest independently to help with your understanding of concepts. All answers should be written in your own words to help with your understanding of concepts.

### Solutions

1. Go to [http://www.chem4kids.com/files/matter\\_solution.html](http://www.chem4kids.com/files/matter_solution.html)
  - a. What is the difference between a homogenous and a heterogeneous mixture?
  - b. What is a solution?
  - c. Give an example of a solution that does not involve a solid being dissolved in a liquid. Explain how this is a solution.
  - d. Define the following terms:
    - Solute
    - Solvent
    - Solubility

### Defining Acids & Bases

2. Go to <https://www.youtube.com/watch?v=ZiokgPOaZ1E>
  - a. How does the Arrhenius theory define acids and bases?
  - b. What is the main problem with the Arrhenius definitions?
  - c. How does the Bronsted-Lowry theory define acids and bases?

### Naming Acids

3. Go to <https://www.youtube.com/watch?v=LoSLKUN7FQ4>
  - a. What is a binary acid?
  - b. How do you name a binary acid?
  - c. Give the formula or name for the following binary acids:

Name	Formula
Hydrochloric Acid	
Hydrobromic Acid	
	HF
	H <sub>2</sub> S

- d. What is an oxyacid?
- e. How do you name an oxyacid?
- f. Give the formula or name for the following oxyacid's:

Name	Formula
Nitric Acid	
Sulfuric Acid	
Phosphoric Acid	
	HNO <sub>3</sub>
	HClO <sub>4</sub>

# Introduction To Acids Bases A Webquest Answer Key

**Ludwig Frederick Audrieth**

## Introduction To Acids Bases A Webquest Answer Key:

*The Lewis Acid-base Concepts* William B. Jensen, 1980      Acids, Bases and Salts MCQ (Multiple Choice Questions)  
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*Reactions of Acids and Bases in Analytical Chemistry* Adam Hulanicki, 1987

Acids and Bases Brian G. Cox, 2013-01-31

Acids and bases are ubiquitous in chemistry. Our understanding of them however is dominated by their behaviour in water. Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates. This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents. The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented. Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution. Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial. Experimental methods for determining dissociation constants are described in detail. In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran. Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change. Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described.

**Proton Chemistry** Robert A. Richardson, 1976 Alberta Authorized Resource for grade 12 ca 1980 1994

*Acids, Bases and Non-aqueous Systems* Ludwig Frederick Audrieth, 1949

**Acids,**

**Bases, and the Chemistry of the Covalent Bond** Calvin Anthony VanderWerf, 1964

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*Holt Science Spectrum Physical Science*

*Chapter 9 Resource File: Acids, Bases, and Salts* Holt Rinehart & Winston, 2008-01-01

Fundamentals of Acids, Bases,

Buffers and Their Application to Biochemical Systems Ralph Scorpio, 2000-08-01

**Hard and Soft Acids and Bases** Ralph

G. Pearson, 1973

Acids, Bases and Salts, 1968

*Acid-bases in Analytical Chemistry* Izaak Maurits Kolthoff, 1966

1997, 1997, 1997, 1997, 1997

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