

Activity B: Methane reaction	Get the Gizmo ready: <ul style="list-style-type: none"> • Select C + H₂ becomes CH₄. 	
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Goal: Predict the results of a chemical reaction.

1. **Count:** CH₄ is the chemical formula for methane. In order to produce a single molecule of methane, how many carbon (C) atoms are needed? **1** Hydrogen atoms? **4**

2. **Predict:** Set the number of C atoms to 14 and the number of H₂ molecules to 10.

A. How many total hydrogen atoms are there? **20**

B. How many methane molecules could form from these reactants? **5**

C. After the atoms react, which reactant will be left over? **The carbon**

D. Which reactant will be the limiting reactant? **Hydrogen**

E. Click Play and wait until Reaction complete is shown. What happened?

Each hydrogen molecule formed a connection with carbon (4 hydrogen molecules per carbon). There are extra carbon atoms that did not link because there are additional carbon atoms.

3. **Test:** Click Reset. Set the number of C atoms to seven.

A. How many hydrogen molecules (H₂) will be needed for there to be no excess reactants? **14** Use the slider to set the number of H₂ molecules to this value.

B. How many methane molecules will be formed? **7**

C. Click Play. What happened?

Since all of the hydrogen molecules linked to the carbon, there were no extra reactants.

4. **Solve:** For each **chemical equation** below, write the number of product molecules that will form from the reaction. Then, circle the **limiting reactant**. (Note: The **coefficients** in front of the reactants indicate the number of reactant molecules or atoms present.)

A. 5C + 6O₂ **5** CO₂

B. 4Na + 8Cl₂ **4** NaCl

Full Version Student Exploration Limiting Reactants

Answer Key Gizmo

R Barnett

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