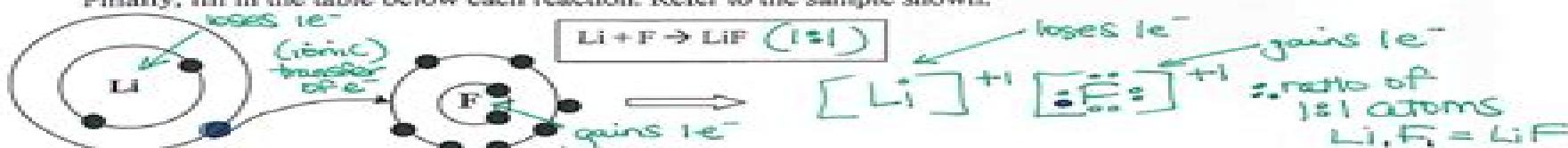


## Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1ve	transfers 1ve to F	Li <sup>+</sup>
F	7ve	accepts 1ve from Li	F <sup>-</sup>

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) Li + Cl $\rightarrow$ LiCl	Li	1ve	donates 1e <sup>-</sup>	Li <sup>+</sup>
	Cl	7ve	accepts 1e <sup>-</sup>	Cl <sup>-</sup>
2) Ca + O $\rightarrow$ CaO	Ca	2ve	donates 2e <sup>-</sup>	Ca <sup>+</sup>
	O	6ve	accepts 2e <sup>-</sup>	O <sup>2-</sup>
3) Be + F $\rightarrow$ BeF <sub>2</sub>	Be	2ve	donates 2e <sup>-</sup> to each F.	Be <sup>+</sup>
	F	7ve	each F accepts 1ve each	F <sup>-</sup>
4) Mg + S $\rightarrow$ MgS	Mg	2ve	donates 2e <sup>-</sup> to S	Mg <sup>+</sup>
	S	6ve	accepts 2e <sup>-</sup>	S <sup>2-</sup>
5) K + F $\rightarrow$ KF	K	1ve	donates 1e <sup>-</sup>	K <sup>+</sup>
	F	7ve	accepts 1e <sup>-</sup>	F <sup>-</sup>

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