

CHAPTER TESTS ANSWER KEY

Matter and Change, pp. 1-4

2.	s.	c.
3.	s.	c.
4.	s.	a.
9. mass and volume		
10. group or family		
11. homogeneous		
12. products		
13. metal		
14. solid state		
15. chemical properties		
16. atom		
17. compound		
18. nonmetal		
19. mixture		
20. pure substance		
21. basic		
22. metalloid		
23. reactants: carbon and oxygen; product: carbon dioxide		
24. reactant: mercury(II) oxide; products: mercury and oxygen		
25. f		
26. g		
27. c		
28. a		
29. b		
30. i		
31. d		
32. h		
33. e		
34. physical change		
35. chemical change		
36. physical change		
37. physical change		
38. chemical change		
39. physical change		
40. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow past one another. In a gas, particles are at greater distances from one another, compared to the particles of liquids and solids.		
41. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.		
42. Although energy can be absorbed or released in a		

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.

44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

Measurements and Calculations, pp. 5-8

2.	s.	c.
3.	c	s. c.
4.	c	t. a.
9. time		
10. mass		
11. density		
12. energy		
13. length or distance		
14. volume		
15. area		
16. qualitative		
17. quantitative		
18. qualitative		
19. quantitative		
20. 3.00×10^5 km/s		
21. three		
22. 0.026 g or 2.6×10^{-2} g		
23. 2.5×10^{11} L		
24. quantity		
25. derived units		
26. conversion factor		
27. 0.0432 kg or 4.32×10^{-2} kg		
28. 5400 mL or 5.4×10^3 mL		
29. 300. K		
30. 1.05×10^7 J		
31. 3.51×10^3 cm		
32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.		
33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.		
34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.		
35. The graph of two quantities that are directly proportional is a straight line through the origin.		

Glencoe Chemistry Matter And Change Answer Key

Chapter 16

Judd E. Hollander

Glencoe Chemistry Matter And Change Answer Key Chapter 16:

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Chemistry 7e Chapter 16 Dickson,1999-08 *Virginia Glencoe Chemistry: Matter and Change* ,2013 *Chemistry* Laurel Dingrando,McGraw-Hill Staff,Glencoe/McGraw-Hill,2001-08-01

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